

**CHEMISTRY 100 — Useful Equations Sheet**

$$\text{number of moles} = \frac{\text{mass}}{\text{molar mass}} = \frac{\text{mass}}{M_r}$$

$$\text{concentration (molarity)} = \frac{\text{number of moles}}{\text{volume}}$$

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

**For dilutions:**  $C_1V_1 = C_2V_2$

**In acid-base equilibria:**

$$\text{pH} = -\log_{10} [\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log_{10} [\text{OH}^-]$$

$$\text{pKa} = -\log_{10} (\text{Ka})$$

$$[\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14}$$

$$\text{pH} + \text{pOH} = 14$$

**For the acid dissociation equilibrium:**



**For the base dissociation equilibrium:**



**Ideal gas law:**  $PV = nRT$

**Gas constant** =  $0.082057 \text{ L.atm.K}^{-1}.\text{mol}^{-1}$  OR  $8.314 \text{ J.mol}^{-1}\text{K}^{-1}$

**Combined gas law:**  $P_1V_1 = P_2V_2$  OR  $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

**Freezing point depression:**  $\Delta T_f = i K_f m_{\text{solute}}$

where  $i$  is the Van't Hoff factor and  $m_{\text{solute}}$  is the molality of the solute (in  $\text{kg.mol}^{-1}$ ) and  $K_f$  is the freezing point depression constant.

$$\text{concentration (molality)} = \frac{\text{moles of solute (mol)}}{\text{mass of solvent (kg)}}$$

$$\text{mass percent (m/m)} = \frac{\text{mass of solute (g)}}{\text{mass of solute (g)} + \text{mass of solvent (g)}} \times 100\%$$

**First Order Integrated Rate Law:**  $\ln (C_t) = -kt + \ln(C_0)$

Where  $C_t$  is the concentration at time  $t$ , and  $C_0$  is the initial concentration with rate constant  $k$ .

**Half-life:**  $t_{1/2} = (\ln 2)/k$

**Avogadro's number** =  $6.022 \times 10^{23} \text{ mol}^{-1}$

**Temperature in Kelvin** = Temperature in Celsius + 273.15

# Pkgacp Chem 1006

**A Gutmann**



**Pkgacp Chem 1006:**

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