CHEM 1211 Fall 2021

Exam 4

These are the equations that will be provided on exam 4.

Equations

 $w = -P\Delta V$ $\Delta E = q + w$ $q = mC_{\nu}\Delta T$ $q_{\nu\nu} = C_{\nu\nu}\Delta T$ $\Delta H^* = \Sigma n\Delta H^*(products) - \Sigma n\Delta H^*(reactants)$ E=hv $c=\lambda v$ $\Delta E = -2.18 \times 10^{-18} J(\frac{1}{n_f^2} - \frac{1}{n_f^2})$ $\lambda = h/mv$ PV = nRT

Constants

1 L atm = 101.325 J STP = 273 K and 1 atm R = 0.08206 L·atm/mol·K = 8.314 J/mol·K = 8.314 kPa·L/mol·K Avogadro's number: 6.022 × 10²⁸ J = 1 kg·m²-s⁻² C_0 (water) = 4.184 J/g*C Mass of an electron: 9.109 x 10^{-31} kg Mass of a proton: 1.673 x 10^{-27} kg $c = 3.00 \times 10^8$ m/s $h = 6.626 \times 10^{-34}$ Js $R_{tot} = 2.18 \times 10^{-18}$ J

Temperature

0 K = -273.15°C K = °C + 273.15 °C = (5/9)(°F - 32)

Volume

 $1 \text{ mL} = 1 \text{ cm}^3$

Pressure.

1 pascal = 1 N/m³ = 1 kg/m·s² 1 atm = 101.325 kPa = 760 mm Hg = 760 torr = 14.70 lb/in² 1 bar = 1 × 10³ Pa (exactly)

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