

# CER Practice

## CER Practice 1

**Question:** Does a chemical change occur if you mix sodium chloride solution and sucrose solution?

### Materials:

- sodium chloride solution (5 g of sodium chloride dissolved in 50 ml of water)
- sucrose solution (5 g of glucose dissolved in 50 ml of water)
- graduated cylinder
- balance
- clear glass beaker
- stirring rod
- thermometer
- red food coloring
- blue food coloring

### Procedure:

1. Place 5 g of sodium chloride into 50 ml of water and stir until it is dissolved.
2. Add red food coloring.
3. Place 5 g of sucrose into 50 ml of water and stir until it is dissolved.
4. Add blue food coloring.
5. Place a thermometer in the sodium chloride solution.
6. Pour the glucose solution into the sodium chloride solution.
7. Record your observations and the temperature.

### Data/Observations:

	Reactants		Products
	sodium chloride solution	sucrose solution	sodium chloride solution with sucrose solution
			clear, transparent
			25°C
			purple
			liquid

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and Print**

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## CER Practice 2

**Question:** Does a chemical change occur if you mix magnesium sulfate and sodium carbonate?

### Materials:

- magnesium sulfate solution (10 g of magnesium sulfate dissolved in 100 ml of water)
- sodium carbonate solution (5 g of sodium carbonate dissolved in 50 ml of water)
- test tube
- two droppers
- graduated cylinder
- balance
- clear plastic cup
- stirring rod
- thermometer

### Procedure:

1. Place 10 g of magnesium sulfate into 100 ml of water and stir until it is dissolved.
2. Place 5 g of sodium carbonate into 50 ml of water and stir until it is dissolved.
3. Place a thermometer into the magnesium sulfate solution and record the starting temperature.
4. Place the sodium carbonate solution into the magnesium sulfate solution.
5. Record your observations and the temperature.

### Data/Observations:

Appearance
Temperature
Color
State of Matter
Other

✓ **State your Claim**

✓ **Back up with Evidence**

✓ **Support with Reasoning**

# Chemical Reactions

# Pkgacp Cer Chemistry 1

**Brooks/Cole**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the Brooks/Cole text.

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