

Laboratory 7: Coordination Complexes: Determination of K using spectrophotometry

LAB SECTION: CHEM 1002 - 423

A. Introduction/Goals of Experiment

In chemistry, chemical equilibrium is the fundamental key concept. It is the state where the rates of both forward and reverse reactions are the same and stay constant. For a reaction:



when studying the equilibrium of chemical reaction, it is important to determine the equilibrium constant K, the ratio of the concentration of the products to the concentrations of the reactants.

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

The value of chemical constant may be determined from experimental data when the concentrations of both reactants and products are known.

In this experiment, we will determine the value of K for the reaction of the formation of a coordinate complex that involves the reaction of ion (III) and thiocyanate.



The goals of this experiment are to study the coordination complex formation and to use Beer's Law to determine the solution's concentration. In order to determine the amount of the complex $\text{Fe}(\text{SCN})^{2+}$, a spectrophotometer will be needed in the experiment. Beer's Law describes the relationship between the absorbance of the solution and the concentration of $\text{Fe}(\text{SCN})^{2+}$:

$$A = sb [\text{Fe}(\text{SCN})^{2+}]$$

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