

CHEM 1412 Atomic structure summary

Introduction

Atomic structure refers to the composition and arrangement of particles within an atom. Atoms are the basic building blocks of matter, and understanding their structure is crucial to understanding the properties of elements and the formation of compounds.

Components of Atoms

Atoms consist of three main subatomic particles:

- Protons: Positively charged particles found in the nucleus. They define the atomic number (Z) of an element.
- Neutrons: Neutral particles also located in the nucleus. The number of neutrons varies among isotopes of a given element.
- Electrons: Negatively charged particles that orbit the nucleus in specific energy levels called electron shells or orbitals. Their arrangement dictates an element's chemical properties.

Atomic Mass and Atomic Number

The atomic mass (A) is the sum of protons and neutrons in the nucleus. The atomic number (Z) represents the number of protons in an atom, which is unique for each element. The number of electrons in a neutral atom is equal to the number of protons. Isotopes are atoms of the same element that have different numbers of neutrons.

Atomic Models

Several atomic models have been proposed over time to explain the structure and behavior of atoms:

- Dalton's Atomic Theory (1803): Atoms are indivisible particles, with each element having unique atomic properties.
- J.J. Thomson's Plum Pudding Model (1897): Electrons are embedded in a positive sphere, like "plums" in a "pudding."
- Rutherford's Nuclear Model (1911): Most of an atom's mass is concentrated in a dense, positively charged nucleus, with electrons orbiting around it.
- Bohr's Model (1913): Electrons orbit the nucleus in specific energy levels at fixed distances.
- Quantum Mechanical Model (1926): Electrons occupy orbitals defined by probability distributions rather than fixed paths, with specific energy levels and shapes.

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Wolfgang Guggemos



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