

CHEM 10062: General Chemistry 1 Laboratory - Lab Exam 2 Review

Lab 10 – Calorimetry, Enthalpy and Hess's Law

Example for the determination of an enthalpy change of a reaction: ΔH_{rxn}

➤ Hint: enthalpy changes are usually referred to i.e. mols of reactant (kJ/mol)



We use 0.052 g of Mg and 50.0 mL of a 1 M HCl solution ($d = 1.02 \text{ g/mL}$) and measure a temperature increase of 4.7°C .

The specific heat of the resulting solution is $3.97 \text{ J/(g}\cdot^\circ\text{C)}$.

The calorimeter constant was just determined.

Calculate the enthalpy change per mole of Mg.

Lab Exam 2 – Topics Covered

- Lab 1: Laboratory Safety and General Lab procedures
- Lab 2: Spectroscopy, Absorbance, and Wavelength
- Lab 3: Beer's Law, Absorbance & Concentration
- Lab 4: Limiting Reagents
- Lab 5: Precipitates & Water Quality
- Lab 6: Titration of Acids in a Beverage
- Lab 7: Calorimetry, Enthalpy and Hess's Law

• Study Guide and practice questions available on Canvas
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• Review your graded Lab Reports

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