

Multiple-choice Questions, 1-14 (Canvas), are 2 points each and Questions 16-19 are 6 points each. (Total Points 52)
* For full credit on problems 15-18 show all of your work, including equations, calculations AND units*

- ___1) Which of the following compounds will be more soluble in acidic solution than in pure water?
A) PbCl_2 B) FeBr_2 C) $\text{Ca}_3(\text{PO}_4)_2$ D) NaI
E) None of the above will be more soluble in acidic solution.
- ___2) Determine the molar solubility of $\text{Fe}(\text{OH})_2$ in pure water. K_{sp} for $\text{Fe}(\text{OH})_2 = 4.87 \times 10^{-17}$.
A) $2.44 \times 10^{-17} \text{ M}$
B) $1.62 \times 10^{-17} \text{ M}$
C) $4.03 \times 10^{-9} \text{ M}$
D) $3.65 \times 10^{-6} \text{ M}$
E) $2.30 \times 10^{-6} \text{ M}$
- ___3) A solution containing CaCl_2 is mixed with a solution of $\text{Li}_2\text{C}_2\text{O}_4$ to form a solution that is $2.1 \times 10^{-5} \text{ M}$ in calcium ion and $4.75 \times 10^{-5} \text{ M}$ in oxalate ion. What will happen once these solutions are mixed? $K_{sp}(\text{CaC}_2\text{O}_4) = 2.3 \times 10^{-9}$.
A) A precipitate will form since $Q > K_{sp}$ for calcium oxalate.
B) Nothing will happen since both calcium chloride and lithium oxalate are soluble compounds.
C) Nothing will happen since calcium oxalate is extremely soluble.
D) Nothing will happen since $K_{sp} > Q$ for all possible precipitants.
E) There is not enough information to determine.
- ___4) The molar solubility of Ag_2S is $1.26 \times 10^{-16} \text{ M}$ in pure water. Calculate the K_{sp} for Ag_2S .
A) 1.59×10^{-32}
B) 1.12×10^{-8}
C) 6.81×10^{-63}
D) 3.78×10^{-12}
E) 8.00×10^{-48}
- ___5) A solution contains $3.8 \times 10^{-2} \text{ M}$ in Al^{3+} and 0.29 M in F^- at equilibrium. If the K_f for AlF_6^{3-} is 7×10^{19} , what is the concentration of the AlF_6^{3-} ion at equilibrium?
A) $1.52 \times 10^{15} \text{ M}$
B) $3.1 \times 10^{-22} \text{ M}$
C) $9.1 \times 10^{19} \text{ M}$
D) $1.9 \times 10^{-21} \text{ M}$
E) $9.6 \times 10^{15} \text{ M}$
- ___6) Which of the following processes have a $\Delta S > 0$?
A) $2 \text{NH}_3(\text{g}) + \text{CO}_2(\text{g}) \rightarrow \text{NH}_2\text{CONH}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$ B) lithium fluoride forms from its elements
C) $2 \text{HBr}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \text{Br}_2(\text{l})$ D) sodium chloride dissolves in pure water
E) All of the above processes have a $\Delta S > 0$.

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Table of Contents Pkg Acp Chemistry 115

1. Understanding the eBook Pkg Acp Chemistry 115
 - The Rise of Digital Reading Pkg Acp Chemistry 115
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Chemistry 115
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Chemistry 115
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Chemistry 115
 - Personalized Recommendations
 - Pkg Acp Chemistry 115 User Reviews and Ratings
 - Pkg Acp Chemistry 115 and Bestseller Lists
5. Accessing Pkg Acp Chemistry 115 Free and Paid eBooks
 - Pkg Acp Chemistry 115 Public Domain eBooks
 - Pkg Acp Chemistry 115 eBook Subscription Services
 - Pkg Acp Chemistry 115 Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Pkg Acp Chemistry 115 Compatibility with Devices
 - Pkg Acp Chemistry 115 Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Pkg Acp Chemistry 115
 - Highlighting and Note-Taking Pkg Acp Chemistry 115
 - Interactive Elements Pkg Acp Chemistry 115
8. Staying Engaged with Pkg Acp Chemistry 115
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkg Acp Chemistry 115
9. Balancing eBooks and Physical Books Pkg Acp Chemistry 115
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Chemistry 115
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pkg Acp Chemistry 115
 - Setting Reading Goals Pkg Acp Chemistry 115
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pkg Acp Chemistry 115
 - Fact-Checking eBook Content of Pkg Acp Chemistry 115
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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