

EXPERIMENT XI

ESTERIFICATION - The Preparation of Methyl Benzoate

REFERENCE: Introduction to Organic Laboratory Techniques, 3rd Ed., Pavia, Lampman and Kriz
Vacuum Distillation, pp. 558-562, 590-592

OBJECTIVES: The student should be able to:

1. Write up a step by step procedure for the reaction using information on the handout sheet, laboratory manuals and The Handbook of Chemistry and Physics.
2. Extract the reaction product by use of the properties of polarity and acidity.
3. Recover the unreacted benzoic acid.
4. Distill the product by vacuum distillation.
5. Calculate the amounts of reactants given the grams of product and the expected percent yield.
6. Identify the final product from physical properties, IR and NMR spectra.

main reaction equation

**REAGENTS:**

Reactants: Benzoic acid; methanol

Catalyst: conc. H_2SO_4 , 4 ml

Calculate the amounts of the reactants required to produce 21.8 grams of methyl benzoate assuming a 80% yield.

Use this theoretical amount of benzoic acid.

Use a 518% excess of methanol.

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Table of Contents Pkg Acp Organic Chemistry 243 Lab

1. Understanding the eBook Pkg Acp Organic Chemistry 243 Lab
 - The Rise of Digital Reading Pkg Acp Organic Chemistry 243 Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Organic Chemistry 243 Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Organic Chemistry 243 Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Organic Chemistry 243 Lab
 - Personalized Recommendations
 - Pkg Acp Organic Chemistry 243 Lab User Reviews and Ratings
 - Pkg Acp Organic Chemistry 243 Lab and Bestseller Lists
5. Accessing Pkg Acp Organic Chemistry 243 Lab Free and Paid eBooks
 - Pkg Acp Organic Chemistry 243 Lab Public Domain eBooks
 - Pkg Acp Organic Chemistry 243 Lab eBook Subscription Services
 - Pkg Acp Organic Chemistry 243 Lab Budget-Friendly Options
6. Navigating Pkg Acp Organic Chemistry 243 Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkg Acp Organic Chemistry 243 Lab Compatibility with Devices
 - Pkg Acp Organic Chemistry 243 Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkg Acp Organic Chemistry 243 Lab
 - Highlighting and Note-Taking Pkg Acp Organic Chemistry 243 Lab
 - Interactive Elements Pkg Acp Organic Chemistry 243 Lab
8. Staying Engaged with Pkg Acp Organic Chemistry 243 Lab

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Pkg Acp Organic Chemistry 243 Lab
- 9. Balancing eBooks and Physical Books Pkg Acp Organic Chemistry 243 Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Organic Chemistry 243 Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkg Acp Organic Chemistry 243 Lab
 - Setting Reading Goals Pkg Acp Organic Chemistry 243 Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkg Acp Organic Chemistry 243 Lab
 - Fact-Checking eBook Content of Pkg Acp Organic Chemistry 243 Lab
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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