Balance elements that appear in only 1 compound on each side (C and H) C2H3N2 + N2O4 -+ N2 +4H2O + 2CO2 (4 Oxygens) (8 Oxygens) Balance elements that appear in more than 1 compound on each side (O) C2H3N2 + 2N2O4 → N2 +4H2O + 2CO2 (2 Nitrogens) (6 Nitrogens) Balance elements that appear in elemental form on either side last (N) C2H3N2 + 2N2O4 → 3N2 +4H2O + 2CO2 Au253 + H2 → Au + H25 (1 Sulfur) (3 Sulfur) Au253 + H2 → Au + 3H25 (2 Gold) (1 Gold) $Au2S3 + H2 \rightarrow 2Au + 3H2S$ (2 Hydrogen) (6 Hydrogen) Au253 + 3H2 → 2Au + 3H25 If they do not add up correctly... use a fraction C3H6 + O2 → CO2 + H2O (3 Carbon) (1 Carbon) C3H6 + O2 → 3CO2 + H2O ((6 Hydrogen) (2 Hydrogen) C3H6 + O2 → 3CO2 + H2O (2 Oxygen) (7 Oxygens) C3H6 + 4.5O2 → 3CO2 + 3H2O Multiply by 2 to make everything a whole number 2C3H6 + 9O2 → 6CO2 + 6H2O Polyatomics are considered separately CaCO3 + NaOH → Ca(OH)2 + Na2CO3 (1 Sodium) (2 Sodium) CaCO3 + 2NaOH → Ca(OH)2 + Na2CO3 Mols and masses of reactants and products using stoichiometric factors (mol ratios) $2H2 + O2 \rightarrow 2H2O$ This means that two moles of H2 reacts with one mol of O2 to generate 2 moles of H2O We can use this to create ratios that can be used in calculations

grams of H2 do we need? 20 moles of H2 2 mol H2O/ 1 mol of O2 is a conversion factor specific to the reaction, called the mol ratio (10.00 moles O2)(2 mol H2O/1 mol O2) = 20.00 mol H2O produced (20.00 mol H2O)(18.02 g/mol) = 360.4 g H2O

How many moles of H2O can be obtained by 10.00 moles O2 (with excess H2) How many

How many H2 is required to react with all 10.00 mol of O2?

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