

Write balanced net ionic equations for the reactions with a correct state for every species.

16. Write the balanced formula for the following half-reactions according to your notes:

Half-Reaction Species	Complete Reaction
permanganate ion	$\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$
nitrate ion	$\text{NO}_3^- \rightarrow \text{NO}$
iodate ion	$\text{IO}_3^- \rightarrow \text{I}_2$
chromate ion	$\text{CrO}_4^{2-} \rightarrow \text{Cr}^{3+}$
perchlorate ion	$\text{ClO}_4^- \rightarrow \text{Cl}^-$
arsenate ion	$\text{AsO}_4^{3-} \rightarrow \text{As}_2\text{S}_3$

16. Write the following half-reactions according to your notes:

Half-Reaction Species	Half-Reaction Species
I_2	$\text{I}^- \rightarrow \text{IO}_3^-$
NO_3^-	$\text{NO}_3^- \rightarrow \text{NO}_2$
Ag^+	$\text{Ag}^+ \rightarrow \text{Ag}$
ClO_2	$\text{ClO}_2 \rightarrow \text{ClO}_2^-$
Mg^{2+}	$\text{Mg}^{2+} \rightarrow \text{Mg}$
H_2O_2	$\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O}$

16. Write the balanced formula for the following half-reactions according to your notes:

Half-Reaction Species	Complete Reaction
chromium(VI) ion	$\text{CrO}_4^{2-} \rightarrow \text{Cr}^{3+}$
gold (III) ion	$\text{Au}^{3+} \rightarrow \text{Au}$
nitrate ion	$\text{NO}_3^- \rightarrow \text{NO}$
iron (II) ion	$\text{Fe}^{2+} \rightarrow \text{Fe}^{3+}$
iodate ion	$\text{IO}_3^- \rightarrow \text{I}_2$
manganate(VI) ion	$\text{MnO}_4^{2-} \rightarrow \text{Mn}^{2+}$

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