

**Unit 1 Nuclear Chemistry**

Size, structure and stability of the nucleus - n/p ratio, packing fraction, mass defect and binding energy - nuclear fission and fusion, atom bombs -hydrogen bomb – radioactivity, alpha, beta particles and gamma radiation - Soddy-Fajan displacement law, half and average life period - Geiger-Muller Counter and Wilson Cloud Chamber, applications of radioactivity - in medicine, agriculture, carbon and fossil dating - isotopes, isobars, isotones, isodiapheres and nuclear isomers - natural and artificial radioactivity, artificial transmutation of elements, induced radioactivity, preparation of transuranic elements, Q values, nuclear coulombic barrier .

**Unit 2 Solid State**

Crystalline and amorphous solids, isotropy and anisotropy, elements of symmetry in crystal systems indices - Miller indices, space lattice and unit cell, Bravais lattices, the seven crystal systems and their Bravais lattices, X-ray diffraction - Bragg's equation and experimental methods (powder method and rotating crystal technique), types of crystals - molecular, covalent, metallic and ionic crystals - close packing of spheres – hexagonal, cubic and body centered cubic packing, interstices in packing - types of crystals – molecular, covalent, metallic crystals - defects in crystals – stoichiometric, non-stoichiometric, extrinsic and intrinsic defects.

**Unit 3 Liquid state**

Properties of liquids-viscosity, surface tension, capillary action, evaporation, vapour pressure, boiling point and distillation, heat transfer involving liquids

**Unit 4 Acids, Bases and Non-aqueous solvents**

Concepts of acids and bases – hard and soft acids and bases - Pearson's concept, HSAB principle and its application - basis for hard - hard and soft - soft interactions - non-aqueous solvents - general characteristics of non-aqueous solvent - melting point, boiling point, latent heat of fusion and vaporization, and dielectric constant - reactions such as complex formation, redox, precipitation and acid base type in non-aqueous solvents like liquid ammonia, liquid  $\text{SO}_2$  and liquid  $\text{HF}$ .

**Unit 5 Water Technology**

Soft and hard water – Hardness – units of hardness – alkalinity - dissolved oxygen – water for various types of industries – treatment of water by ion exchange process - boiler feed water – boiler compounds – internal and external conditioning - water for drinking - municipal water treatment – desalination by RO and electro dialysis.

**TEXTBOOKS:**

1. Marion Clyde Day Jr, Joel Selbin, [Harry H Sisler](#), '*Theoretical Inorganic Chemistry*', LLC, 2012
2. F. A. Cotton and G. Wilkinson, '*Advanced Inorganic Chemistry*', 5<sup>th</sup> edition, John Wiley and Sons, New York, 1987
3. B. R. Puri, L. R. Sharma, Kalia, '*Principles of Inorganic Chemistry*', Vishal Publishing Co., 2008

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1. H. S. Arniker, '*Essentials of Nuclear Chemistry*', 4<sup>th</sup> edition, New Age International Publishers, 2005.

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## **Table of Contents Pkg Acp Cer General Chemistry Ii Ch 11**

1. Understanding the eBook Pkg Acp Cer General Chemistry Ii Ch 11
  - The Rise of Digital Reading Pkg Acp Cer General Chemistry Ii Ch 11
  - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Cer General Chemistry Ii Ch 11
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pkg Acp Cer General Chemistry Ii Ch 11
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Cer General Chemistry Ii Ch 11
  - Personalized Recommendations
  - Pkg Acp Cer General Chemistry Ii Ch 11 User Reviews and Ratings
  - Pkg Acp Cer General Chemistry Ii Ch 11 and Bestseller Lists
5. Accessing Pkg Acp Cer General Chemistry Ii Ch 11 Free and Paid eBooks
  - Pkg Acp Cer General Chemistry Ii Ch 11 Public Domain eBooks
  - Pkg Acp Cer General Chemistry Ii Ch 11 eBook Subscription Services
  - Pkg Acp Cer General Chemistry Ii Ch 11 Budget-Friendly Options

- 
6. Navigating Pkg Acp Cer General Chemistry Ii Ch 11 eBook Formats
    - ePub, PDF, MOBI, and More
    - Pkg Acp Cer General Chemistry Ii Ch 11 Compatibility with Devices
    - Pkg Acp Cer General Chemistry Ii Ch 11 Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Pkg Acp Cer General Chemistry Ii Ch 11
    - Highlighting and Note-Taking Pkg Acp Cer General Chemistry Ii Ch 11
    - Interactive Elements Pkg Acp Cer General Chemistry Ii Ch 11
  8. Staying Engaged with Pkg Acp Cer General Chemistry Ii Ch 11
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Pkg Acp Cer General Chemistry Ii Ch 11
  9. Balancing eBooks and Physical Books Pkg Acp Cer General Chemistry Ii Ch 11
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Pkg Acp Cer General Chemistry Ii Ch 11
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Pkg Acp Cer General Chemistry Ii Ch 11
    - Setting Reading Goals Pkg Acp Cer General Chemistry Ii Ch 11
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Pkg Acp Cer General Chemistry Ii Ch 11
    - Fact-Checking eBook Content of Pkg Acp Cer General Chemistry Ii Ch 11
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements

- 
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