CHEM 132 Principles of Chemistry Lab II Montgomery

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Principles of Chemistry Lab II Montgomery College, Rockville

Acids and Bases, pH, Buffers and Hydrolysis Introduction

Acids and Bases

Aqueous solutions of acids and bases are recognized as "acidie" or "basic" because they contain

appreciable concentrations of either hydronium (H3O+) or hydroxide (OH-) ions. Hydronium ions are

produced from the reaction of covalent molecules like HCI with water.

H3O+ (aq) + Cl- (aq) HCl (g) + H2O (l) Some bases are ionic compounds that dissolve in water or react with it to produce aqueous hydroxide

ions, for example:

NaOH (s) Na+ (aq) + OH- (aq) BaO (s) + H2O (l) Ba2+ (aq) + 2OH-(aq) Other bases such as ammonia, NH3, and related amines produce hydroxide ions when dissolved in water

by accepting a proton from the solvent H2O.

NH4+ (aq) + OH- (aq) NH3 (g) + H2O (l) Acids and bases are classified as either strong or weak. Strong acids or bases are those that are

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Ian Pickup

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