

Stoichiometry practice problems

13. 13.2 g of iron(III) alum were dissolved in water and reduced to an iron(II) ion solution by zinc and dilute sulfuric acid. The mixture was filtered and the filtrate and washings were made up to 500 cm³ in a standard volumetric flask. If 20.0 cm³ of this solution required 26.5 cm³ of 0.0100 mol dm⁻³ KMnO₄ for oxidation:

- (a) give the ionic equation for the reduction of iron(III) ions by zinc metal.
(b) Calculate the percentage by mass of iron in iron alum.

$$n = CV$$

$$n = \frac{m}{M_r}$$

$$2 : 3$$

Mole ratios

Problems In Stoichiometry

J Ma



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