

# Noble Gas Compounds

Table 22.1 Properties of Xenon Compounds

Compound	Oxidation State of Xe	Melting Point (°C)	$\Delta H_f^\circ$ (kJ/mol) <sup>a</sup>
XeF <sub>2</sub>	+2	129	-109(g)
XeF <sub>4</sub>	+4	117	-218(g)
XeF <sub>6</sub>	+6	49	-298(g)
XeOF <sub>4</sub>	+6	-41 to -28	+146(l)
XeO <sub>3</sub>	+6	— <sup>b</sup>	+402(s)
XeO <sub>3</sub> F <sub>2</sub>	+6	31	+145(s)
XeO <sub>4</sub>	+8	— <sup>c</sup>	—

<sup>a</sup>At 25 °C, for the compound in the state indicated.

<sup>b</sup>A solid; decomposes at 40 °C.

<sup>c</sup>A solid; decomposes at -40 °C.

- Discovered in 1962
- Xe forms most of these compounds
- Lower ionization energy and expanded valence shell lead to reactions.
- KrF<sub>2</sub> (decomposes at -10 °C) and a few other Kr compounds (unstable above -40 °C) also can be made.

# Noble Gas Compounds

**Lingsheng Yao**



## **Noble Gas Compounds:**

*The Chemistry of the Noble Gases* Cedric L. Chernick, 1967      **Noble-gas Chemistry** John H. Holloway, 1968      Noble Gases and Their Compounds G. J. Moody, John David Ronald Thomas, 1964      Noble Gas Compounds Herbert Harvey Hyman, 1963-12      *Noble-Gas Compounds* Herbert H. Hyman, 1963      **Noble gases**, 1970      The Noble Gases Adam Furgang, 2010-01-15 This book explains the characteristics of the noble gases helium neon argon krypton xenon and radon including where they are found and how they are used by humans      *Noble Gas Compounds a Bibliography 1962-1976* Warren E. Falconer, Neil Bartlett, 1978-02      **The Noble Gases** Howard H. Claassen, 1966      *Neon and the Noble Gases* Nigel Saunders, 2003 This series uses a common or well known element to look at the groups of the periodic table and to show the similarities and differences between elements It uses full colour illustration of the periodic table and shows the chemical symbol for each element in place alongside its neighbours Chemical formulae for common compounds are also shown Information boxes and tables contain listings of facts and figures Chemical reactions are interpreted as word equations and timelines chart the history and discovery of the elements atoms and why they don't even react with themselves It guides you through the discovery and isolation of each gas and recent experiments by scientists to make noble gas compounds It also explains why helium balloons float neon lights glow and argon atmospheres protect metals from air

**Noble-Gas Compounds. [Papers Read at the 1963 Conference on Noble-Gas Compounds Held at Argonne National Laboratory.] Edited with Introductions by Herbert H. Hyman. [With Plates.].** Argonne National Laboratory. Conference on Noble-Gas Compounds, Herbert Hyman HYMAN, 1963      *Noble Gas Chemistry* Felice Grandinetti, 2018-06-22 Authored by one of the world's leading experts in the chemistry of lighter noble gases this comprehensive monograph fills the need for an up to date review of the diverse experimental techniques and theoretical methods currently in practice After reviewing the experiments breaking the paradigm of non reactive noble gases the physico chemical background is introduced Besides the emphasis on gas phase reactions the author presents other relevant systems such as chemistry in the bulk phase under high pressure and cold matrices The discussion of gas phase chemistry of the noble gases covers neutral and ionic compounds diatomic molecules complexes with small molecules and metal compounds up to large clusters      Noble Gas Compounds Jeffery E. Fisher, 1966      *The Formation of Noble Gas Compounds* Leland Cullen Allen, 1965      **Halogens and Noble Gases, Second Edition** Monica Halka, Brian Nordstrom, 2019-12-01 In spite of their adjacency in the periodic table halogens and nonmetals have very different properties Halogens are among the most chemically reactive elements in the periodic table exhibiting a diverse chemistry in terms of the large numbers of compounds they can form On the other hand noble gases are the least chemically reactive elements In fact before the 1960s chemists referred to these elements as inert gases because it was believed that they exhibited no chemistry whatsoever Providing the basics of these elements including their role in history and some of the important scientists involved in their discovery this newly updated full color resource

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