

Alliances

- Loose alliances began to form at the same time that rivalries over colonies and trade grew
- **Alliance** → An agreement or friendship between two or more parties, made in order to advance common goals and to secure common interests

Loose Alliances

J Spring



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was produced i what mass of iodine was produced

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160 0

g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ $26.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield $\frac{\text{theoretical yield}}{\text{actual yield}} \times 100$ based on this definition we would expect a percent yield

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

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 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 \times 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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