

Key

1. Assign an oxidation number to each of the following indicated elements:

Pb(NO ₃) ₂	Cl ₂	ClF	LiCl	Cr(NO ₃) ₃	CoBr ₂	N ₂ O ₅
Pb <u>+4</u>	Cl <u>0</u>	Cl <u>+1</u>	Li <u>+1</u>	Cr <u>+3</u>	Co <u>+2</u>	N <u>+5</u>
N <u>+3</u>		F <u>-1</u>	Cl <u>-1</u>	N <u>+5</u>	Br <u>-1</u>	
MnO ₂	Mn ₂ O ₃	KMnO ₄	CF ₂ Cl ₂	Pd(OH) ₂	CsClO ₃	
Mn <u>+4</u>	Mn <u>+3</u>	Mn <u>+7</u>	C <u>+4</u> F <u>-1</u>	Pd <u>+2</u> O <u>-2</u>	Cs <u>+1</u> Cl <u>+5</u>	
O <u>-2</u>		K <u>+1</u>	Cl <u>-1</u>	H <u>+1</u>		

2. Consider the following oxidation-reduction reaction.



Write the two half-reactions. Label each half-reaction as an oxidation or a reduction.



Which species is oxidized?

C in CO

What is the reducing agent?

C in CO

3. In the smelting of iron from iron oxide according to the equation



- (a) C in CO (b) O in Fe₂O₃ (c) C in CO₂ (d) Fe metal (e) Fe in Fe₂O₃

4. Identify and justify the Bronsted-Lowry base in the reaction: $\text{NH}_3 + \text{H-Br} \rightarrow \text{NH}_4^+ + \text{Br}^-$

- (a) NH₃ because it is a proton donor (b) HBr because it is a proton donor (c) NH₃ because it is a proton acceptor (d) HBr because it is a proton acceptor

5. A solution is prepared by dissolving 1.28 g of NaNO₃ in enough water to make 500 mL of solution. What is the concentration of the solution in molarity? (The unit of molarity is moles / liter given the symbol M)

$$\begin{aligned} & 1.28 \text{ g NaNO}_3 \\ & \div 14.0067 \\ & = 15.5598 \times 3 \\ & = 46.6794 \end{aligned}$$

$$\frac{1.28 \text{ g NaNO}_3}{84.9947 \text{ g NaNO}_3} \times \frac{1 \text{ mol NaNO}_3}{84.9947 \text{ g NaNO}_3} = 0.0151 \text{ mol NaNO}_3$$

$$\frac{0.0151 \text{ mol}}{0.5 \text{ L}} = 0.0302 \text{ M}$$

6. What is the weight (in grams) of NaNO₃ in a 0.500 L bottle of 2.15 M NaNO₃?

$$\frac{2.15 \text{ mol NaNO}_3}{1 \text{ L}} \times 0.5 \text{ L} \times \frac{84.9947 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 91.4 \text{ g NaNO}_3$$

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