



📄 Key Points and Links

▾ Section 1.2: Describe the chemical properties of matter

📄 Matter

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📄 Classifying Matter

▸ Section 1.3: Identify minerals and explain their formation

▸ Section 1.4: Explain the formation and implications of the rock cycle



around the outside of the nucleus. Until the twentieth century, it was thought that atoms were small clumps of protons and electrons bound together in a solid chunk. However, a series of deceptively simple experiments conducted from 1908 to 1913 showed that atoms are mostly empty space. All elements heavier than hydrogen (H) contain another type of particle in the nucleus, called neutrons. Neutrons are like nuclear glue; they are responsible for holding all of the protons in the nucleus of an atom.

The mass of an atom is determined by the number of its protons, neutrons, and electrons. You can figure out the mass of an atom by adding up the total mass from the protons, neutrons, and electrons it contains. The mass of a proton is about 1837 times larger than the mass of an electron, and the mass of a neutron is similar to the mass of a proton. Therefore, the mass of an

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**Guillaume Favre**



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