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Love Stinks



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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{CuNO}_3$$

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ $260.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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C is the limiting reagent 4 33 g of H₂ are left over 13

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yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is

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