

DICCIONARIO

ÉTICA CRISTIANA

Y TEOLOGÍA PASTORAL

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Pastoral Spanish 2nd Edition Rev

Michael Seilmaier



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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N_2 what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x 1 mol Al x 2 mol Al_2O_3 x 101.964 g Al_2O_3 302.3 g Al_2O_3 26.982 g Al 4 mol Al 1 mol Al_2O_3 c determine the percent yield $\frac{\text{yield actual} \times 100}{260.0 \text{ g}} \times \frac{100}{86.01} = \frac{\text{theoretical}}{302.3 \text{ g}}$

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 $80\text{ g I}_2\text{O}_5 \times \frac{1\text{ mol I}_2\text{O}_5}{1\text{ mol I}_2\text{O}_5} \times \frac{1\text{ mol I}_2}{1\text{ mol I}_2\text{O}_5} \times \frac{253.8\text{ g I}_2}{1\text{ mol I}_2} = 28\text{ g CO} \times \frac{1\text{ mol CO}}{28\text{ g CO}} \times \frac{1\text{ mol I}_2}{5\text{ mol CO}} \times \frac{253.8\text{ g I}_2}{1\text{ mol I}_2} = 25.38\text{ g I}_2$

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