



		ACTIONS

Performance Expectation	Use the periodic table as a model to predict the relative properties of elements based on the patterns of electrons in the outermost energy level and the composition of the nucleus of atoms. Physical Science: Examples of properties that could be predicted from patterns could include metals, nonmetals, metalloids, number of valence electrons, types of bonds formed, or atomic mass. Emphasis is on main group elements. Chemistry: Examples of properties that could be predicted from patterns could include reactivity of metals, types of bonds formed, numbers of bonds formed, atomic radius, atomic mass, or reactions with oxygen. Emphasis is on main group elements and qualitative understanding of the relative trends of ionization energy and electronegativity.		
Science & Engineering Practices	Disciplinary Core Ideas	Crosscutting Concepts	
	STRUCTURE AND PROPERTIES OF MATTER Each atom has a charged substructure consisting of a nucleus, which is made of protons and neutrons, surrounded by electrons. [HS.PS1A.a) The periodic table orders elements horizontally by the number of protons in the atom's nucleus and places those with similar chemical properties in columns. The repeating patterns of this table reflect patterns of outer electron states.[HS.PS1A.b] TYPES OF INTERACTIONS Attraction and repulsion between electric charges at the atomic scale explain the structure, properties, and transformations of matter, as well as the contact forces between material objects.[HS.PS2B.c)	PATTERNS Different patterns may be observed at each of the scales at which a system is studied and can provide evidence for causality in explanations of phenomena.	



Science Interactions Performance Assessment Course 4

Robert W. Avakian

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