

Table 2 Colors of Several Acid-Base Indicators

Indicator	Color in more acidic range	Transition range (pH)	Color in more basic range
A	pink	1–2	yellow
B	pink	3–4	yellow
C	yellow	4–6	blue
D	red	5–8	blue
E	colorless	8–10	pink
F	colorless	10–12	yellow

26. Which of the following indicators can be used to best detect the equivalence point for both Titrations 1 and 4?

- A) Indicator B
- B) Indicator C
- C) Indicator E
- D) Indicator F

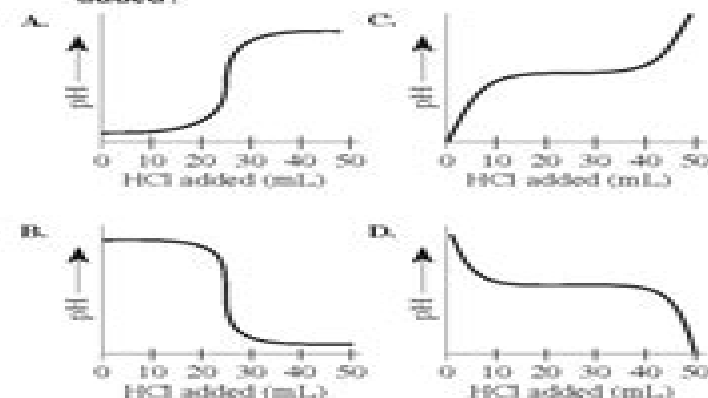
27. Which of the following expressions gives an approximate value of the equilibrium constant (K_a) of the acid in Titration 5?

- A) $\log([\text{H}_3\text{O}^+]^2 \times [\text{HOAc}])$
- B) $[\text{H}_3\text{O}^+]^2 \times [\text{HOAc}]$
- C) $\log([\text{H}_3\text{O}^+]^2/[\text{HOAc}])$
- D) $\frac{[\text{H}_3\text{O}^+]^2}{[\text{HOAc}]}$

28. The concentrations of H_3O^+ and OH^- ions are equal after approximately how much of NaOH is added during Titration 3?

- A) 4 mL
- B) 7 mL
- C) 25 mL
- D) 50 mL

29. Assume that Titration 2 is modified so that HCl is titrated into NaOH. Which of the following graphs best shows how the pH changes as HCl is added?



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