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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^{-}$ is

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web google classroom for the following redox equation $\text{MnO}_4^{-} + \text{aq} \rightarrow \text{SO}_4^{2-} + \text{aq}$ $\text{H}_2\text{SO}_4 + \text{aq}$ what will be the stoichiometric coefficient of H_2SO_4 if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-}$ Fe^{2+} Cr^{3+} Fe^{3+} solution 1 balanced half reactions 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ Fe^{2+} Fe^{3+} e 2 equalize the electrons 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ 6Fe^{2+} 6Fe^{3+} 6e^- multiplied by a factor of 6

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