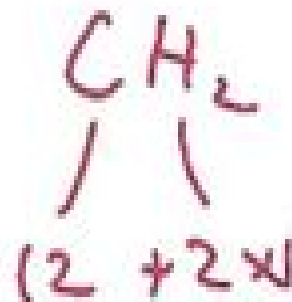


2.) Cyclopropane is a compound of carbon and hydrogen that is used as a general anesthetic. When 1.00 g of this substance was burned completely in oxygen, 3.14 g of CO_2 and 1.29 g of H_2O were produced. What is the empirical formula of the compound? The molar mass of cyclopropane is about 42 g/mole.

$$\frac{3.14}{44} = 0.0714$$

$$\frac{1.29}{18} = 0.0716$$



0.1433 moles of H.

$$\begin{array}{r} \text{C} : \text{H} \\ \hline 0.0714 : 0.1433 \\ \hline 0.0714 \quad 0.0714 \\ \hline 1 : 2 \end{array}$$

Selected Combustion Problems Ii Transpor

LL Leslie



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