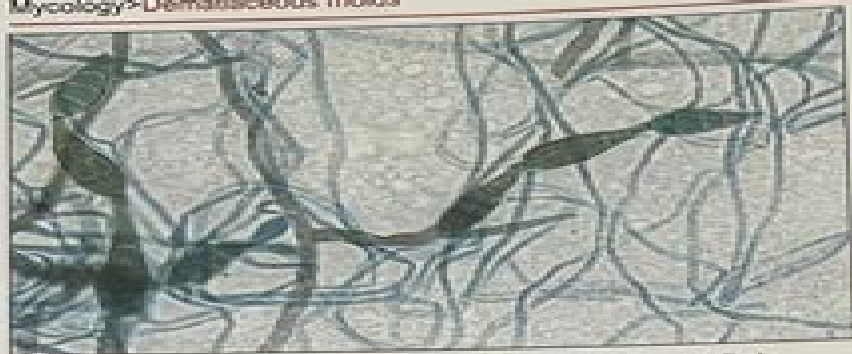
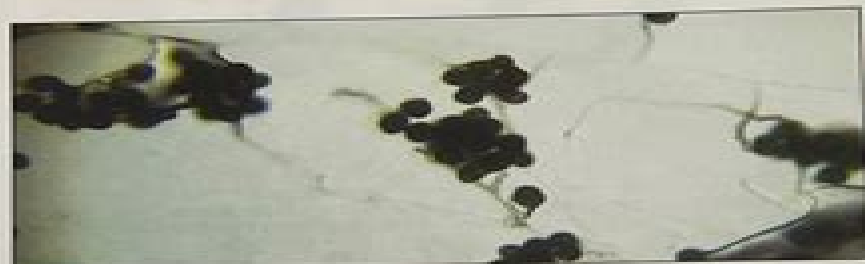


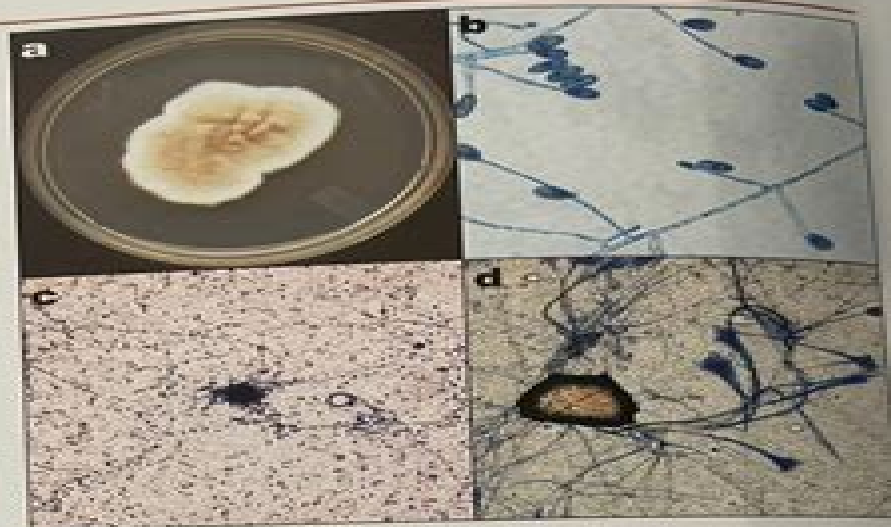
Mycology > Dematiaceous molds



16.99 *Alternaria* species form chains of conidia with transverse & longitudinal (uniform) septations; one end of each conidium is blunt & the other pointed



16.99 *Ulocladium*: oval conidia borne singly on geniculate conidiophores



16.100 *Pseudallescheria boydii* & *Scoedosporium boydii* complex
a forms a mold colony with a light brown melanized surface
b oval, truncated, melanized microconidia with nonmelanized hyphae
c in the sexual state, the mold produces large dark cleistothecia
d the alternative asexual form, *Graphium*, characterized by thick mats of long conidiophores stuck together side by side, resembling the bristles of a broom

□ *Alternaria*

- Considered a contaminant but may cause disease in tissue, nails, sinuses
- Colony surface is brown or greenish black & reverse is black
- Chains of club-shaped conidia with transverse & longitudinal septations in the conidia & have alternating blunt & pointed ends **16.98**

□ *Ulocladium*

- Considered a contaminant
- Colony is dark brown to black with a black reverse
- Smooth or rough oval conidia borne singly on bent (geniculate) conidiophores **16.99**

■ Slow growing dematiaceous molds

□ *Pseudallescheria boydii* & *Scoedosporium boydii* complex

- *P boydii* is sexual state & *S boydii* is asexual state
- Common cause of mycetoma

- Light gray or brown surface **16.100a**, owing to melanization of its oval, truncated microconidia **16.100b**
- Hyphae are hyaline (nonmelanized), resulting in a light reverse
- In its sexual state, produces large (50-200 µm), dark cleistothecia **16.100c**
- Asexual state does not form cleistothecia but produces conidiophores that bear conidia singly or in small groups; they resemble mold phase of *Blastomyces dermatitidis*
- Alternative asexual form produces *Graphium* type conidiation; it is characterized by thick mats of long conidiophores stuck together side by side, resembling the bristles of a broom **16.100d**
- *P boydii* & *S boydii* complex resistant to amphotericin B, but usually susceptible to triazoles such as voriconazole & posaconazole

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant k is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of k

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^{-}$

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web the chemical equation used was $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product was determined to be $K_{\text{sp}} \text{Ca(OH)}_2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant ksp chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} \text{Ca}_3(\text{PO}_4)_2 = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated Ca(OH)_2 soultion with a st

standard hcl solution and find the ksp or the oh concentration was found and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the ksp of calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq for this experiment the titration of calcium hydroxide with hydrochloric acid was investigated determining the ksp of calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq for this experiment the titration of calcium

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for Ca(OH)_2 was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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