

Chapitre 2 Analyser un système chimique par des méthodes physiques

1) Analyse par mesure d'absorbance

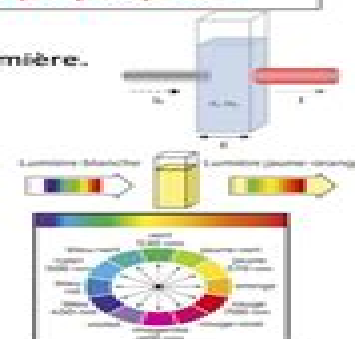
L'absorbance est une grandeur physique caractérisant la capacité d'une solution à absorber la lumière.

Loi de Beer-Lambert :

A une longueur d'onde donnée, l'absorbance A_λ d'une espèce chimique en solution diluée est proportionnelle à la concentration molaire C de cette espèce et à l'épaisseur l de la solution traversée :

$$A_\lambda = \epsilon_\lambda \times l \times C = k \times C$$

absorbance de la solution (sans unité) coefficient d'extinction molaire ($\text{L.mol}^{-1}.\text{cm}^{-1}$)
Longueur de la cuve (cm) concentration molaire de l'espèce chimique (mol.L^{-1})



La loi de Beer-Lambert n'est valable que pour des solutions diluées dont l'absorbance est inférieure à 2.

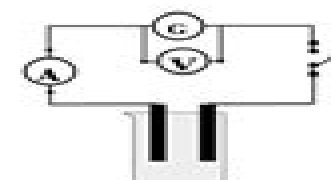
2) Analyse par mesure de conductance

activité expérimentale page 43

a) Les solutions conductrices

Pour qu'une solution conduise le courant électrique, il faut qu'elle contienne **des ions**.

Dans une solution ionique, les **cations** se déplacent dans le sens conventionnel du courant, alors que les **anions** se déplacent dans le sens



b) Conductance d'une solution ionique

La **conductance** d'une portion de solution comprise entre 2 électrodes permet d'estimer le caractère conducteur d'une solution. C'est l'inverse de la résistance.

$$G : \text{conductance en Siemens (S)} \quad R : \text{résistance en Siemens } (\Omega) \quad G = \frac{1}{R} = \frac{I}{U}$$

I : intensité du courant (A)
 U : tension entre les électrodes (V)

c) Conductivité d'une solution

On constate expérimentalement, que la conductance est proportionnelle à :

la surface des électrodes

l'inverse de la longueur entre les électrodes.

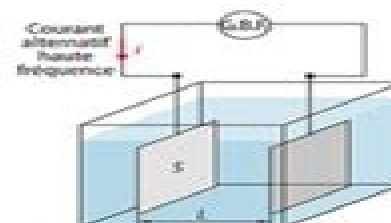
- Le coefficient de proportionnalité s'appelle la **conductivité de la solution** :

$$G = \sigma \times \frac{S}{L}$$

conductivité de la solution (S.m^{-1}) surface des électrodes (m^2)
distance entre les électrodes (m)

On constate expérimentalement que la conductance G augmente avec la **température** de la solution ainsi qu'avec la **concentration** en soluté apporté de la solution et dépend de l'état de la surface des électrodes (propre, lisse, ...).

La **conductivité** d'une solution aqueuse est une grandeur qui représente la capacité de cette solution à conduire le courant électrique. Elle se note σ et s'exprime en **siemens par mètre**



$$\sigma = G \times \frac{L}{S} = k \times G$$

conductivité (S) constante de cellule (m^{-1}) conductance (S.m^{-1})

Methodes Physiques Appliquees A La Chimie

L Towner



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