

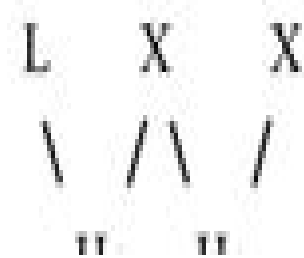
Name	Molecular formula	Structure weight	Molecular solubility	Water constant	Henry's law
Mercury	Hg	–	200.59	61 µg/L	$8.7 \times 10^{-3}$
Mercury(I) chloride	Hg <sub>2</sub> Cl <sub>2</sub>	Cl-Hg-Hg-Cl	472.09	4.0 mg/L	–
Mercury(I) bromide	Hg <sub>2</sub> Br <sub>2</sub>	Br-Hg-Hg-Br	560.99	22 µg/L	–
Mercury(I) iodide	Hg <sub>2</sub> I <sub>2</sub>	I-Hg-Hg-I	654.99	0.24 µg/L	–
Mercury(I) oxylate	Hg <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	[2Hg] <sup>2+</sup> [O <sub>2</sub> -C-C-O] <sup>2-</sup>	489.20	3.5 mg/L	–
Mercury(II) chloride	HgCl <sub>2</sub>	Cl-Hg-Cl	271.52	73 g/L	$7.1 \times 10^{-6}$
Mercury(II) bromide	HgBr <sub>2</sub>	Br-Hg-Br	360.40	6.1 g/L	–
Mercury(II) iodide	HgI <sub>2</sub>	I-Hg-I	454.40	0.14 µg/L	–
Mercury(II) oxalate	HgC <sub>2</sub> O <sub>4</sub>	[Hg] <sup>2+</sup> [O <sub>2</sub> -C-C-O] <sup>2-</sup>	288.61	0.4 g/L	$1.4 \times 10^{-9}$
Mercury(II) oxide	HgO	Hg-O	216.59	53 mg/L	$7.1 \times 10^{-7}$
Methylmercury chloride	CH <sub>3</sub> HgCl	CH <sub>3</sub> -Hg-Cl	251.10	0.1 g/L	$4.7 \times 10^{-7}$
Dimethylmercury	C <sub>2</sub> H <sub>5</sub> Hg	CH <sub>3</sub> -Hg-CH <sub>3</sub>	230.66	1 g/L	$7.6 \times 10^{-3}$

**Notes:** Henry's law constant, expressed as the ratio of the partial pressure above the solution to the solution concentration ( $K_{\text{H}} = P_{\text{aq}}/c_{\text{aq}}$ ), at 25°C, is given in atm·m<sup>3</sup>/mol. Data from: Clever et al.<sup>8</sup> Dry and Gledhill,<sup>10</sup> National Research Council,<sup>12</sup> Iverfeldt and Lindqvist,<sup>13</sup> and Sillman et al.<sup>14</sup>

stable complexes of this type is the basicity of the ligand.<sup>18</sup>

Nitrogen-containing ligands of higher basicity (pK<sub>a</sub> >5.2)

result in the disproportionation of Hg<sub>2</sub><sup>2+</sup> due to the much



# Metals In Mercury Solubility Data Series Vol 25

**I. Lambert, H.L. Clever**



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