



Oscar De La Renta

Léo Pomar



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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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web 11 use avogadro s law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22 4 l at stp v 1 n 1 v 2 n 2 0 0896 l n 22 4 l 1 mole n 0 0040 moles 12 experimental molar mass g mole

of butane 0.23 g 0.0040 moles 58 g mole 13 the molecular formula of butane is C_4H_{10}

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web the pressure of the gas the pressure of the gas was calculated using dalton's law of partial pressures by subtracting the pressure of the water vapor from the total pressure which in this case was the atmospheric pressure 760 mm Hg the pressure of the water vapor was found by using the temperature of the water

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