

A NEWMARKET PARENTING GUIDE

THE READY-TO-READ, READY-TO-COUNT HANDBOOK

HOW TO BEST PREPARE YOUR CHILD FOR SCHOOL



**A STEP-BY-STEP PROGRAM
FOR TEACHING PRESCHOOLERS
LETTERS AND NUMBERS**

TERESA SAVAGE

ILLUSTRATED WITH DRAWINGS AND CARTOONS • SECOND EDITION

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EW Minium



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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S x 18g H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ grams O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ grams O}_2} = 0.666 \text{ moles O}_2$$

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g Al}} \times \frac{2 \text{ mole Al}_2\text{O}_3}{4 \text{ mole Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 150.1 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{126.9 \text{ g Al}_2\text{O}_3}{150.1 \text{ g Al}_2\text{O}_3} \times 100 = 84.5\%$$

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO to determine the mass of iodine I_2 which could be produced.
 $80 \text{ g } \text{I}_2\text{O}_5 \times \frac{1 \text{ mol } \text{I}_2\text{O}_5}{1 \text{ mol } \text{I}_2} \times \frac{1}{333.8 \text{ g } \text{I}_2\text{O}_5} \times 28 \text{ g } \text{CO} \times \frac{1 \text{ mol } \text{CO}}{28 \text{ g } \text{CO}} = 1 \text{ mol } \text{I}_2$

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 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$ label 3.7.3