

PROBLEMS:

- 1) Calculate the quantities of Lime & Soda required to soften 5000 litres of water containing the following salts:

$\text{MgCl}_2 = 15.5 \text{ ppm}$; $\text{Ca}(\text{HCO}_3)_2 = 32.5 \text{ ppm}$,

$\text{CaSO}_4 = 22.4 \text{ ppm}$; $\text{Mg}(\text{HCO}_3)_2 = 14.6 \text{ ppm}$

$\text{NaCl} = 50 \text{ ppm}$

Soln: Calculation of Calcium Carbonate Equivalents:

Hard Salt	Weight (ppm)	M.Wt	$\text{CaCO}_3 \text{ eq.} = \frac{W}{\text{MW}} \times 100$
MgCl_2	15.5	95	$\frac{15.5 \times 100}{95} = 16.31$
$\text{Ca}(\text{HCO}_3)_2$	32.5	162	$\frac{32.5 \times 100}{162} = 20.06$
CaSO_4	22.4	136	$\frac{22.4 \times 100}{136} = 16.47$
$\text{Mg}(\text{HCO}_3)_2$	14.6	146	$\frac{14.6 \times 100}{146} = 10.00$

Rational Use Of Water And Its Treatment In The Chemical Industries

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