

PROBLEMS:

- 1) Calculate the quantities of Lime & Soda required to soften 5000 litres of water containing the following salts:

$\text{MgCl}_2 = 15.5 \text{ ppm}$; $\text{Ca}(\text{HCO}_3)_2 = 32.5 \text{ ppm}$,

$\text{CaSO}_4 = 22.4 \text{ ppm}$; $\text{Mg}(\text{HCO}_3)_2 = 14.6 \text{ ppm}$

$\text{NaCl} = 50 \text{ ppm}$

Soln: Calculation of Calcium Carbonate Equivalents:

Hard Salt	Weight (ppm)	M.Wt	$\text{CaCO}_3 \text{ eq.} = \frac{W}{\text{MW}} \times 100$
MgCl_2	15.5	95	$\frac{15.5 \times 100}{95} = 16.31$
$\text{Ca}(\text{HCO}_3)_2$	32.5	162	$\frac{32.5 \times 100}{162} = 20.06$
CaSO_4	22.4	136	$\frac{22.4 \times 100}{136} = 16.47$
$\text{Mg}(\text{HCO}_3)_2$	14.6	146	$\frac{14.6 \times 100}{146} = 10.00$

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Table of Contents Rational Use Of Water And Its Treatment In The Chemical Industriessales No E91iie3

1. Understanding the eBook Rational Use Of Water And Its Treatment In The Chemical Industries
- 1.1 The Rise of Digital Reading Rational Use Of Water And Its Treatment In The Chemical Industries
- 1.2 Advantages of eBooks Over Traditional Books
2. Identifying Rational Use Of Water And Its Treatment In The Chemical Industries
- 2.1 Exploring Different Genres
- 2.2 Considering Fiction vs. Non-Fiction
- 2.3 Determining Your Reading Goals
3. Choosing the Right eBook Platform
- 3.1 Popular eBook Platforms
- 3.2 Features to Look for in an Rational Use Of Water And Its Treatment In The Chemical Industries
- 3.3 User-Friendly Interface
4. Exploring eBook Recommendations from Rational Use Of Water And Its Treatment In The Chemical Industries
- 4.1 Personalized Recommendations
- 4.2 Rational Use Of Water And Its Treatment In The Chemical Industries User Reviews and Ratings
- 4.3 Rational Use Of Water And Its Treatment In The Chemical Industries and Bestseller Lists
5. Accessing Rational Use Of Water And Its Treatment In The Chemical Industries Free and Paid eBooks
- 5.1 Rational Use Of Water And Its Treatment In The Chemical Industries Public Domain eBooks
- 5.2 Rational Use Of Water And Its Treatment In The Chemical Industries eBook Subscription Services

- Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3 Budget-Friendly Options
- 6. Navigating Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3 eBook Formats
 - ePub, PDF, MOBI, and More
 - Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3 Compatibility with Devices
 - Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Highlighting and Note-Taking Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Interactive Elements Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
- 8. Staying Engaged with Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
- 9. Balancing eBooks and Physical Books Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Setting Reading Goals Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Rational Use Of Water And Its Treatment In The Chemical Industries sales No E91iie3
 - Fact-Checking eBook Content of Rational Use Of Water And Its Treatment In The Chemical Industries sales No

E91iie3

- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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