

CHEMISTRY

10.) Find the mass of an oxygen gas, O_2 molecule.

- A. $53.1 \times 10^{-24} g$ B. $50.4 \times 10^{-30} g$
C. $45.2 \times 10^{-22} g$ D. $48.4 \times 10^{-27} g$

Solution:

$$\text{mass of molecule} = \frac{\text{molar mass}}{\text{Avogadro's number}}$$

$$\text{mass of } O_2 = \frac{2 \times 16 \text{ g/mol}}{6.022 \times 10^{23} \text{ mol}^{-1}}$$

$$\text{mass of } O_2 = 50.4 \times 10^{-30} g$$

20.) A compound used to whiten teeth has the empirical formula HO . The molecular mass of this compound is 34 amu. What is the molecular formula?

- A. H_2O_3 B. H_3O_3
C. H_2O_2 D. H_3O_2

Solution:

$$\frac{g}{\text{mol}} = \text{amu}$$

$$\text{molar mass of } HO = 1 \text{ amu} + 16 \text{ amu}$$

$$\text{molar mass of } HO = 17 \text{ amu}$$

Divide the molecular mass of the compound by the molar mass of the empirical formula.

$$\frac{34 \text{ amu}}{17 \text{ amu}} = 2$$

Multiply the subscripts of the empirical formula by the whole number.



23.) What is the volume of 2M solution that contains 0.5 mol of solute?

- A. 400 mL B. 250 mL
C. 200 mL D. 500 mL

Solution:

$$2M = 2 \text{ molar} = 2 \text{ mol/liter}$$

$$\text{Volume} = \frac{\text{moles of solute}}{\text{molarity}}$$

$$\text{Volume} = \frac{0.5 \text{ mol}}{2 \text{ mol/L}}$$

$$\text{Volume} = 0.25 \text{ Liters} = 250 \text{ mL}$$

29.) Calculate the mass of water required to prepare a 4M aqueous solution of NH_3 using 25 g of NH_3 .

- A. 0.329 kg B. 0.347 kg
C. 0.368 kg D. 0.356 kg

Solution:

Calculate the moles of NH_3 present in 25 g of NH_3 .

$$\text{moles of } NH_3 = \frac{\text{mass of } NH_3}{\text{molar mass of } NH_3}$$

$$\text{moles of } NH_3 = \frac{25 \text{ g}}{\left(14 \frac{g}{\text{mol}}\right) + 3 \left(1 \frac{g}{\text{mol}}\right)}$$

$$\text{moles of } NH_3 = 1.47 \text{ mol}$$

Calculate the volume of a 4M NH_3 solution required to dissolve 25 g of NH_3 .

$$\text{Volume} = \frac{\text{moles of solute}}{\text{molarity}}$$

$$\text{Volume} = \frac{1.47 \text{ mol}}{4 \text{ mol/L}} = 0.368 \text{ L}$$

Calculate the mass of water required to prepare the required volume of NH_3 solution.

$$\text{Density of water} = 1.00 \text{ g/mL} = 1 \text{ kg/L}$$

$$\text{mass of water} = \text{Volume} \times \text{water density}$$

$$\text{mass of water} = 0.368 \text{ L} \times 1.00 \frac{g}{\text{mL}}$$

$$\text{mass of water} = 0.368 \text{ kg}$$

Merril Chemistry Problems And Solutions Manual

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