

# Modern Europe 1815-1899. Period VIII

W. Alison Phillips

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# Modern Europe Vol 1 Modern Europe To 1815

**James Harvey Robinson, James Henry  
Breasted, Charles Austin Beard**



## **Modern Europe Vol 1 Modern Europe To 1815:**

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temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1210 \text{ J}$   
 $550 \text{ J} = \Delta h \cdot q_{\text{water}} = 550 \text{ J}$

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for water  $4.184 \text{ J/g} \cdot \text{C}$  1 what is the specific heat of aluminum if the temperature of a  $28.4 \text{ g}$  sample of aluminum is increased by  $8.1 \text{ C}$  when  $207 \text{ J}$  of heat is added 2

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 $14.7 \text{ J} = 3.14 \Delta H_{\text{rxn}} / n$   $q_{\text{rxn}} = q_{\text{calorimeter}}$   $m \cdot c \cdot \Delta T$

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web calorimetry worksheet 1 if 0 315 moles of hexane  $\text{C}_6\text{H}_{14}$  is combusted in a bomb calorimeter containing 5 65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 55 4 °C the specific heat capacity of water is 4 184 J/g °C  $m = 5 650$  grams  $\text{H}_2\text{O}$  4 184 J/g °C 55 4 °C  $\Delta H = 1310$  kJ

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web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from ice at 34°C to steam at 138°C a sample of water absorbs  $1 41 \times 10^5$  J find the mass of the sample answers 1 2 68  $\times 10^4$  J 2 6 23  $\times 10^4$  J 3 1 71  $\times 10^4$  J 4 1 11  $\times 10^5$  J 5 2 11  $\times 10^5$  J 6

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