

Problem Set on Heat & Thermodynamics

Thermometry

1. The precision & reliability of 'Modern Platinum Resistance Thermometers' are due to the work of Callender & Griffiths who, from a wide range of experiments, found that the resistance of 'Pure Platinum' is very accurately given by $R_t = R_0(1 + \alpha t + \beta t^2)$ where R_t & R_0 are resistances at $t^\circ\text{C}$ & 0°C respectively and α & β are constants of finite dimensions depending upon the material being tested. Also we know that the platinum temperature is calculated from the formula

$$t_p = (R_t - R_0/R_{100} - R_0) \times 100$$

Show that, using the these two relations, a true temperature t can be measured by the formula:

$$t - t_p = \delta \left\{ \left(\frac{t}{100} \right)^2 - \left(\frac{t_p}{100} \right)^2 \right\}$$

where δ is a constant depending upon the purity of the specimen.

2. A platinum resistance thermometer reads 2.57Ω at 0°C , 3.53Ω at 100°C and 2.89Ω in a bath. Calculate the temperature of the bath on Platinum scale & Celsius scale. Assume $\delta = 1.5$.
3. The resistances of a Platinum wire at 0°C , 100°C and at Sulphur point (444.6°C) are found to be 2.56Ω , 3.56Ω & 6.78Ω respectively. The resistance of the wire at a temperature $t^\circ\text{C}$ is given by $R_t = R_0(1 + \alpha t + \beta t^2)$. Calculate the ratio α/β .
4. The platinum temperature corresponding to 50°C on the gas scale is 50.25°C . find the temperature on the platinum scale that corresponds to 150°C on the gas scale.
5. The resistances of a Platinum wire at 0°C , 100°C and at Sulphur point (444.6°C) are found to be 2.56Ω , 3.56Ω & 6.78Ω respectively. Find the true temperature when the resistance of the thermometer is 5.56Ω .
6. The resistance of a carbon resistor obeys the following relation:

$$\alpha + b \log R = \sqrt{\log R/t}$$
 where t is the temperature. The resistance is found to be exactly $1 \text{ K}\Omega$ in a liquid helium cryostat. If $a = -1.16$ & $b = 0.675$, what is the temperature of the cryostat?

Kinetic Theory of Gases

7. Calculate the pressure exerted by a gas having 2.5×10^{19} molecules per c.c. at a temperature of 27°C .
8. Calculate the r.m.s. velocity of hydrogen at NTP. Given that 1 litre hydrogen weighs 0.0897 g at NTP.
9. At what temperature will the r.m.s. velocity of a gas is half of its value at 0°C ?
10. The mean kinetic energy of a molecule at 27°C is $8 \times 10^{-14} \text{ erg}$. Calculate the number of molecules per cc. of an ideal gas at 27°C and at a pressure of 40 mm of Hg.
11. Estimate the size of a He atom, assuming its mean free path at NTP to be $28.5 \times 10^{-6} \text{ cm}$. Given the density of He at NTP is 0.0178 g/Litre and the mass of He atom is $6 \times 10^{-24} \text{ g}$. (Hint: $\lambda = 1/\sqrt{2}n\sigma$)

Problems In Heat And Thermodynamics

Percy Williams Bridgman



Problems In Heat And Thermodynamics:

Worked Problems in Heat, Thermodynamics and Kinetic Theory for Physics Students L. Pincherle, 2013-10-22

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Engineering Applications of Higher Mathematics: Problems on thermodynamics Vladimir Karapetoff,1916 The Collected Works of J. Willard Gibbs ...: Thermodynamics Josiah Willard Gibbs,1928 *Engineering Thermodynamics, Second Edition* ACHUTHAN, M.,2009-08-30 Thermodynamics is the branch of science concerned with the relations between heat and other forms of energy involved in physical and chemical processes This revised edition of the book continues to provide a thorough understanding of the fundamentals and principles of thermodynamics starting with the most elementary ideas of heat and temperature The book also focuses on practical applications of thermodynamic processes and equips students with simple techniques of solving engineering problems The book also provides systematic problem solving methodology a large number of solved examples a number of review questions at the end of each chapter and a fairly large number of unsolved exercises with hints New to This Edition Includes a set of 107 additional problems in Appendix A set in different examinations

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The Scientific Papers of J. Willard Gibbs, Ph. D. Ll. D., Formerly Professor of Mathematical Physics in Yale University: Thermodynamics Josiah Willard Gibbs, 1906

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