

Chemistry 200: General Chemistry I

Practice Problems for Final Exam

1. If 4.67 kJ of heat is applied to a mixture of 10.0 g ice ($\Delta H_{fus} = 6.02$ kJ/mole) and 25.0 g water ($C_s = 4.18$ J/g $\cdot^\circ$ C) initially at 0.0 $^\circ$ C, what is the final temperature of the mixture?

2. Oxygen difluoride reacts with water to form hydrofluoric acid and molecular oxygen. Using Table 9.3 on p. 392 in your textbook, calculate ΔH_{rxn} .

3. Carbon monoxide can be synthesized according to the reaction below. If 22.0 g C and 42.0 g O₂ are sealed in a 20.00 L stainless steel vessel at 800 $^\circ$ C, what are the values of P_{O_2} and P_{CO} after the reaction? Use $R = 0.08206$ L $\cdot$ atm/K $\cdot$ mole.



4. A redox titration using the equation shown below is performed on a 50.0 mL sample of the oxidizing agent. If the titration required 22.1 mL of a 0.125 M solution of the reducing agent, what is the concentration of the oxidizing agent in the 50.0 mL sample?



5. Calculate the difference in energy between the $n = 2$ and $n = 4$ levels in the hydrogen atom. Use $R = 1.097 \times 10^7$ m⁻¹.

6. Draw the molecular orbital diagram for Ne₂⁺. Calculate its bond order. Is Ne₂⁺ stable or unstable?

7. When hydrochloric acid is added to solid potassium sulfite, the products formed are liquid water, gaseous sulfur dioxide, and aqueous potassium chloride. If a student wants to form 100.0 mL of sulfur dioxide under STP conditions, but can only use 20.0 mL of hydrochloric acid, what must the concentration of the hydrochloric acid be? Use $R = 0.08206$ L $\cdot$ atm/K $\cdot$ mole.

8. Diatomic hydrogen is created and used in the series of reactions below. If a researcher wants to produce 10.0 kJ of heat, how many grams of zinc are needed?



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JS Bruner



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