



## General Chemistry formula sheet

Prefix	Symbol	Factor
Giga	G	$10^9$
Mega	M	$10^6$
Kilo	K	$10^3$
centi	C	$10^{-2}$
milli	m	$10^{-3}$
micro	$\mu$	$10^{-6}$
nano	n	$10^{-9}$
$1 \text{ mol} = 6.022 \times 10^{23} \text{ molecules}$ $1 \text{ cm}^3 = 1 \text{ mL}$		

Name	Formula	Name	Formula
Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	Phosphate	$\text{PO}_4^{3-}$
Carbonate	$\text{CO}_3^{2-}$	Ammonium	$\text{NH}_4^+$
Bicarbonate	$\text{HCO}_3^-$	Chlorite	$\text{ClO}_2^-$
Hydroxide	$\text{OH}^-$	Chlorate	$\text{ClO}_3^-$
Nitrite	$\text{NO}_2^-$	Sulfite	$\text{SO}_3^{2-}$
Nitrate	$\text{NO}_3^-$	Sulfate	$\text{SO}_4^{2-}$
Cyanide	$\text{CN}^-$		

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$P_A = X_A \cdot P_{\text{total}}$$

$$c = \lambda \nu$$

$$M = \frac{\text{mol solute}}{\text{L. solution}}$$

$$PV = nRT$$

$$q = mc\Delta T$$

$$\Delta H^\circ_{\text{rxn}} = \sum n_p \Delta H^\circ_f - \sum n_r \Delta H^\circ_f$$

$$FC = \# \text{ of valence } e^- - [\# \text{ of lone pairs} + \# \text{ of bonds}]$$

$$M_1V_1 = M_2V_2$$

$$X_A = \frac{\text{mol } A}{\text{total mol}}$$

$$q = n\Delta H$$

This instructional aid was prepared by the Tallahassee Community College Learning Commons.



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## **Table of Contents Preview Of General Chemistry**

1. Understanding the eBook Preview Of General Chemistry
  - The Rise of Digital Reading Preview Of General Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Preview Of General Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Preview Of General Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Preview Of General Chemistry
  - Personalized Recommendations
  - Preview Of General Chemistry User Reviews and Ratings
  - Preview Of General Chemistry and Bestseller Lists
5. Accessing Preview Of General Chemistry Free and Paid eBooks
  - Preview Of General Chemistry Public Domain eBooks
  - Preview Of General Chemistry eBook Subscription Services
  - Preview Of General Chemistry Budget-Friendly Options

6. Navigating Preview Of General Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Preview Of General Chemistry Compatibility with Devices
  - Preview Of General Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Preview Of General Chemistry
  - Highlighting and Note-Taking Preview Of General Chemistry
  - Interactive Elements Preview Of General Chemistry
8. Staying Engaged with Preview Of General Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Preview Of General Chemistry
9. Balancing eBooks and Physical Books Preview Of General Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Preview Of General Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Preview Of General Chemistry
  - Setting Reading Goals Preview Of General Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Preview Of General Chemistry
  - Fact-Checking eBook Content of Preview Of General Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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