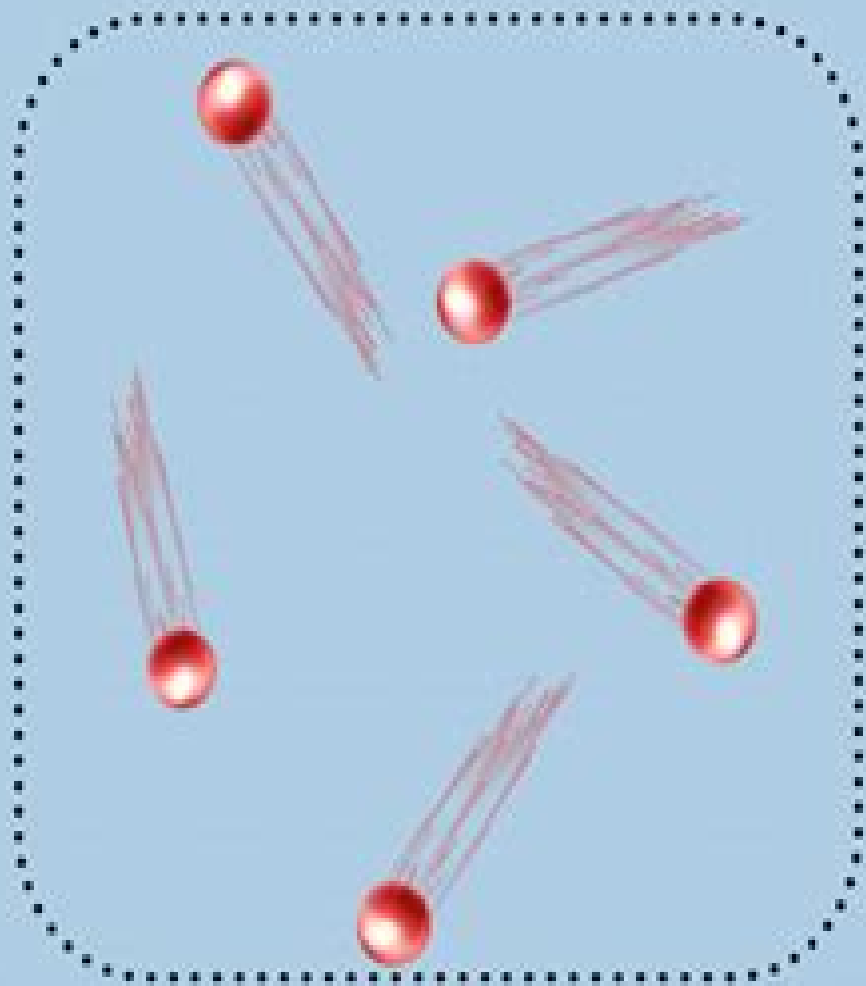


Kinetic Molecular Theory of Gases

- Many gas particles, so statistics applies
- Particles are tiny and spread far apart
- Gas particles are in constant random motion
- No attraction or repulsion between particles
- Collisions are elastic



Molecular Theory Of Gases And Liquids

Joseph Oakland Hirschfelder



Molecular Theory Of Gases And Liquids:

The Molecular Theory of Gases and Liquids Joseph O. Hirschfelder, Charles F. Curtiss, R. Byron Bird, 1964-01-15 An essential cross disciplinary reference for molecular interactions Molecular Theory of Gases and Liquids offers a rigorous comprehensive treatment of molecular characteristics and behaviors in the gaseous and fluid states A unique cross disciplinary approach provides useful insight for students of chemistry chemical engineering fluid dynamics and a variety of related fields with thorough derivations and in depth explanations throughout Appropriate for graduate students and working scientists alike this book details advanced concepts without sacrificing depth of coverage or technical detail **Molecular**

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and experimentalists and engineers who feel intimidated by more formal or rigorous theoretical accounts and textbooks Exercises at the end of each chapter help the reader to apply the basic concepts also to other types of liquid crystal in particular the smectic liquid crystal

Molecular Theory of Gases and Liquids. [By] J. O. Hirschfelder ... Charles F. Curtiss ... R. Byron Bird, etc Joseph Oakland Hirschfelder, Robert Byron Bird, Charles Francis Curtiss, 1954 A Kinetic Theory of Gases and Liquids (Classic Reprint) Richard D. Kleeman, 2017-12

Excerpt from A Kinetic Theory of Gases and Liquids In constructing a general Kinetic Theory the problem that presents itself first for investigation is the dependence of the velocity of translation of a molecule in a substance on its density and temperature It is often assumed that this velocity is the same in the liquid as in the gaseous state at the same temperature It can be shown however that this holds only for each molecule at the instant it passes through a point in the substance at which the forces of the surrounding molecules neutralize each other The total average velocity corresponding to the whole path of a molecule is usually much greater than the foregoing velocity in a liquid and dense gas on account of the effect of the molecular forces of attraction and repulsion About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at www.forgottenbooks.com This book is a reproduction of an important historical work Forgotten Books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy In rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition We do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works

A Kinetic Theory of Gases and Liquids Richard Daniel Kleeman, 2013-09 This historic book may have numerous typos and missing text Purchasers can usually download a free scanned copy of the original book without typos from the publisher Not indexed Not illustrated 1920 edition Excerpt of gas from which the molecule comes and thus increases with decrease of pressure The difference between V_a and V_c therefore also increases with decrease of pressure which increases the value of the ratio T_j i_j according to the foregoing equation This corresponds to a decrease of η and thus the effect of slipping becomes the greater the lower the pressure The dynamical mechanism underlying the result that the viscosity of a gas is independent of its density may be GASEOUS VISCOSITY AND THE TEMPERATURE 127 illustrated by the following considerations A molecule transfers a certain amount of momentum to the gas at the end of each transfer distance on migrating at right angles to the motion of the gas in the direction of the decrease of motion If the concentration of the gas is halved the length of each transfer distance is doubled while the momentum transferred at the end of each transfer distance is also doubled since the velocity gradient of the gas remains the same Since a change in molecular concentration of a gas does not alter the molecular velocities the momentum transferred per second by a molecule moving between two parallel plates of material one of which is at rest while the other moves parallel to itself is in the latter case double that in the former But since the number of molecules per cubic cm available for momentum transference in the former case is half the number in the latter the total

momentum transferred is in each case the same or the viscosity of the gas has not been altered by altering its density The quantity K is a function of the temperature This is shown by the calculated values of K contained in Table X for a number of gases at different temperatures These values are not corrected for

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Table of Contents Molecular Theory Of Gases And Liquids

1. Understanding the eBook Molecular Theory Of Gases And Liquids
 - The Rise of Digital Reading Molecular Theory Of Gases And Liquids
 - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Theory Of Gases And Liquids
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molecular Theory Of Gases And Liquids
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Theory Of Gases And Liquids
 - Personalized Recommendations
 - Molecular Theory Of Gases And Liquids User Reviews and Ratings

- Molecular Theory Of Gases And Liquids and Bestseller Lists
- 5. Accessing Molecular Theory Of Gases And Liquids Free and Paid eBooks
 - Molecular Theory Of Gases And Liquids Public Domain eBooks
 - Molecular Theory Of Gases And Liquids eBook Subscription Services
 - Molecular Theory Of Gases And Liquids Budget-Friendly Options
- 6. Navigating Molecular Theory Of Gases And Liquids eBook Formats
 - ePub, PDF, MOBI, and More
 - Molecular Theory Of Gases And Liquids Compatibility with Devices
 - Molecular Theory Of Gases And Liquids Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molecular Theory Of Gases And Liquids
 - Highlighting and Note-Taking Molecular Theory Of Gases And Liquids
 - Interactive Elements Molecular Theory Of Gases And Liquids
- 8. Staying Engaged with Molecular Theory Of Gases And Liquids
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molecular Theory Of Gases And Liquids
- 9. Balancing eBooks and Physical Books Molecular Theory Of Gases And Liquids
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molecular Theory Of Gases And Liquids
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molecular Theory Of Gases And Liquids
 - Setting Reading Goals Molecular Theory Of Gases And Liquids
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molecular Theory Of Gases And Liquids
 - Fact-Checking eBook Content of Molecular Theory Of Gases And Liquids
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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