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2



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confiance
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Scène écrite
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Anton Keiss



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NE FAITES CONFIANCE A PERSONNE. Brian Lowry,1997-01-01 Avec la troisi me saison de X Files les nouvelles aventures de Mulder et Scully ont fascin des millions de spectateurs Dans ce deuxi me volume le seul guide officiel de la s rie vous permet d approcher la v rit d encore plus pr s avec des documents exclusifs Un r sum complet de tous les pisodes de la troisi me saison dont un pisode in dit en France Les commentaires chaud de Chris Carter David Duchovny et Gillian Anderson Les secrets les anecdotes du tournage de chaque pisode Des photos in dites Les r ponses aux questions les plus fr quemment pos es sur la s rie L interview des sc naristes comment crit on un sc nario d Aux fronti res du r el Et un reportage exceptionnel sur la fabrication compl te d un pisode De la conception par Chris Carter jusqu aux ultimes effets sp ciaux en passant par le tournage suivez le processus qui a donn naissance Anagramme le suspense final de la saison Magie noire mutants spectres vengeurs cafards extraterrestres faiseurs de miracles et conspirations t n breuses D licieusement parano aque et diaboliquement agenc e la s rie la plus envo tante de la t l vision vous entra ne sa suite Ne faites confiance personne

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction

naoh aq c6h4 cooh cook aq c6h4 coona cook aq h2o l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the