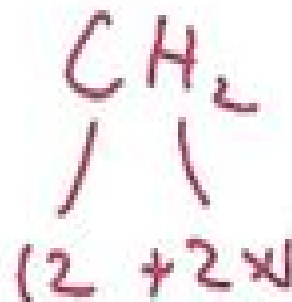


2.) Cyclopropane is a compound of carbon and hydrogen that is used as a general anesthetic. When 1.00 g of this substance was burned completely in oxygen, 3.14 g of CO_2 and 1.29 g of H_2O were produced. What is the empirical formula of the compound? The molar mass of cyclopropane is about 42 g/mole.

$$\frac{3.14}{44} = 0.0714$$

$$\frac{1.29}{18} = 0.0716$$



0.1433 moles of H.

$$\begin{array}{r} \text{C} : \text{H} \\ \hline 0.0714 : 0.1433 \\ \hline 0.0714 \quad 0.0714 \\ \hline 1 : 2 \end{array}$$

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Selected Combustion Problems Ii Transpor**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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