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3 4 mass relationships in chemical equations - Sep 02 2023

web jul 4 2022 to calculate the mass of titanium metal that you can obtain multiply the number of moles of titanium by the molar mass of titanium 47.867 g/mol moles ti mass ti times molar mass ti $4.12 \text{ mol ti} \times 47.867 \text{ g/mol ti} = 197 \text{ g ti}$ notag

chapter 3 mass relationships in chemical reactions - Nov 23 2022

web 41 42 chemistry ch 3 mass relationships in chemical reactions molecular mass of NO_2 atomic mass of n 2 atomic mass of o 14.01 amu $2 \times 16.00 \text{ amu} + 46.01 \text{ amu}$ atomic mass of N_2O_5 2 atomic mass of n 5 atomic mass of o 2 14.01 amu $5 \times 16.00 \text{ amu} + 108.02 \text{ amu}$ example 3 2 exercise 3 3