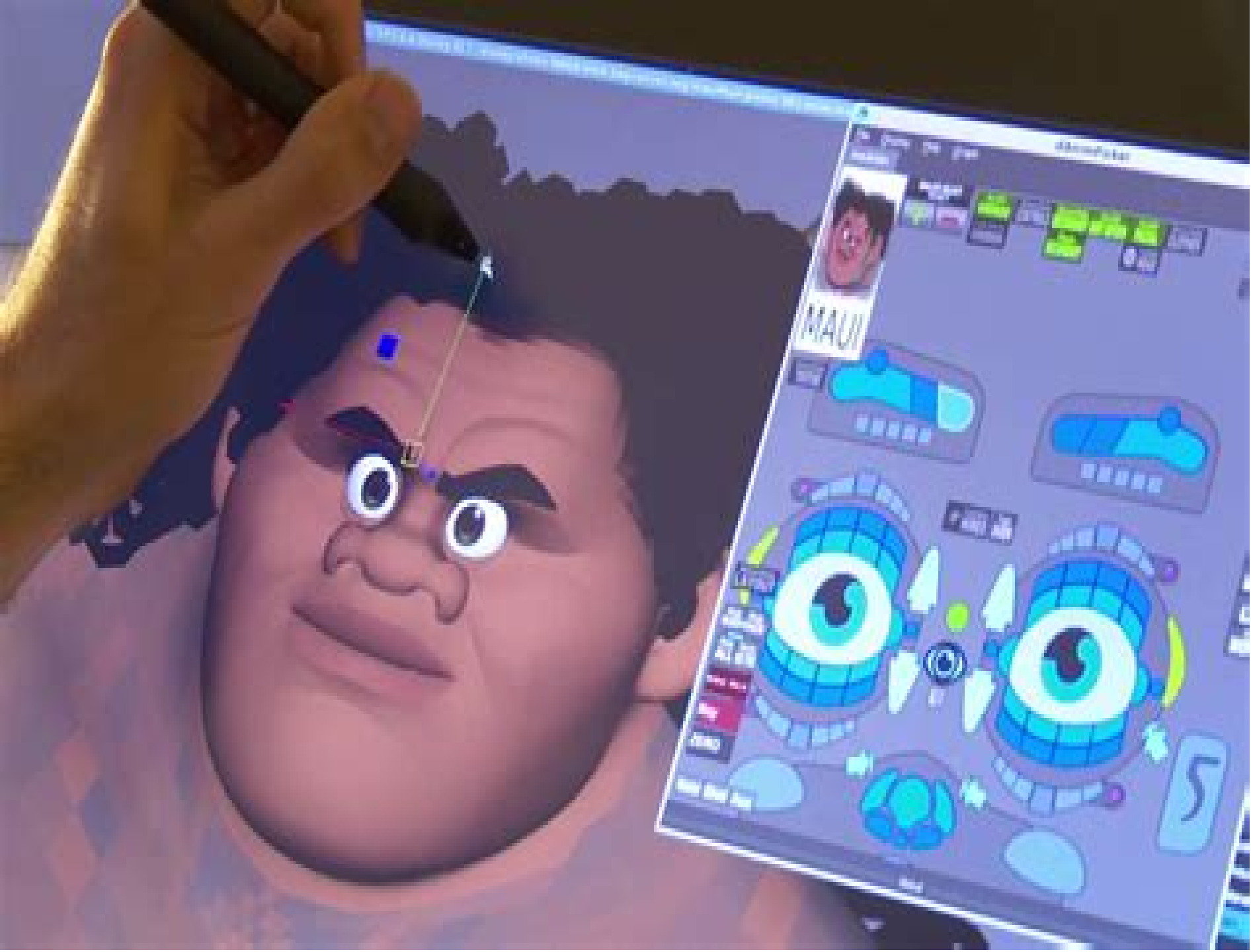




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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution
and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of
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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves
exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ}$
 mol rxn write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is
formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them
salt carbon polypropylene vermiculite water and iron

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specifically instructed to do so 2 never remove chemicals or other materials from the laboratory area 3 do not eat food drink
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