Part I - Multiple Choice (45 pts, 3 pts each).

- 1) Which one of the following is not a state function?
- A) enthalpy H.
- B) heat q
- C) temperature T
- D) pressure P
- E) internal energy U

Use the table of specific heats to answer questions 2 and 3

- 2) Which of the following substances would need the largest amount of heat to change the temperature of 1 gram by 10°C?
- (A) Aluminum, Al
- B) Copper, Cu
- C) Gold, Au
- D) Lead, Pb
- E) Silver, Ag

Substance	Specific Heat
Aluminum	0.900 J / g °C
Copper	0.386 J / g °C
Gold	0.126 J / g °C
Lead	0.128 J / g °C
Silver	0.233 J / g °C

3) Which would need the largest amount of heat to change the temperature of 1 mole by 10°C?

- 3) Which would need the largest associated as a second of the largest as

- Q) Lead, Pb 24, 5
- E) Silver, Ag 26.

4) The enthalpy of formation, ΔH°ε, of NaOH (s) is the same as the ΔH°ενν for which of these reactions?

- A) Na (s) + H₂O (l) → NaOH (s) + ½ H₂ (g)
- B) Na (s) + 35 O2 (g) + 32 H2 (g) → NaOH (s).
- C) Na (s) + ½ H₂O₂ (l) → NaOH (s)
- D) Na* (aq) + OH* (aq) \rightarrow NaOH (s)

Na(5) + O(g) + H(g) -> Na OH

E) N(g) + a + O(g) + H (aq) - NaOH(s)

Incorrect because

O(g) : 1 hot 5 handard

5) The existence of discrete (quantized) energy levels in an atom may be inferred from:

- A) diffraction of electrons by crystals
- B) atomic line spectra
- C) visible spectrum
- D) experiments on the photoelectric effect -> disrovery of photon
- E) X-ray diffraction by crystals

6) Arrange the following in order of increasing energy:

Ultra-violet (UV) radiation, visible radiation, gamma-rays, and infra-red (IR) radiation.

Highest

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