

# HEALTHY AGING WITH NUTRITION



# Nutrition And Aging

**ML Yell**



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web the molar volume will be the volume occupied by one molar mass  $32 \text{ g}$  of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{d}$   $\frac{32 \text{ g}}{1.790 \text{ g/l}} = 17.88 \text{ l}$   $\frac{1.0405 \text{ l}}{1 \text{ mol}} = 1.0405 \text{ l/mol}$  nonnumber

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web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important



point to keep in mind is that  $n$  can be moles of any gas since one mole always contains  $6.02 \times 10^{23}$  molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as  $6.02 \times 10^{23}$  particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always  $6.02 \times 10^{23}$  but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole  $\text{m}^3 \text{mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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web at standard temperature and pressure stp the molar volume  $V_m$  is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass  $M$  by mass density  $\rho$  molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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