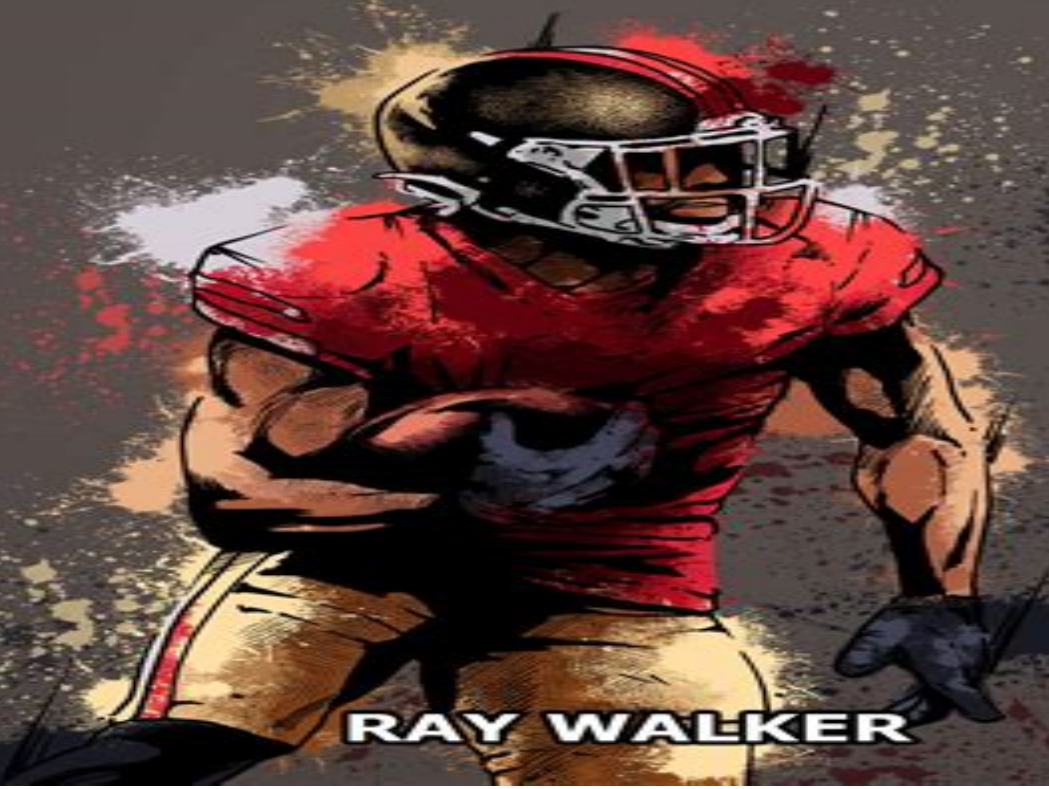


# THE ULTIMATE SAN FRANCISCO 49ERS TRIVIA BOOK

A COLLECTION OF AMAZING TRIVIA QUIZZES AND FUN FACTS  
FOR DIE-HARD 49ERS FANS!



RAY WALKER

# San Francisco 49ers Trivia

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assignment titration curve a titration curve is a plot of the ph of a solution as a function of the amount of base usually in ml

added to an acidic solution or as the amount of acid added to a basic solution the base is added to

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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17.3.2.1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $\text{ph} = \log 0.100 = 1.000$  ph  $\log 0.100 = 1.000$  b titrant volume 12.50 ml

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web expert answer 100% 2 ratings the example is based upon henderson hasselbalch equation the equation gives relation between ph and  $pK_a$  of weak acid the equation is  $\text{ph} = pK_a + \log \frac{[\text{salt}]}{[\text{acid}]}$  at equivalence point all a view the full answer

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web ab acid base titration curves purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the  $pH$  value of a solution vs the amount of titrant added

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web acid base titration lab report jessica chung chem 267 lab date completed 25 september 2019 date turned in 9 october 2019 figure 1 acid base titration curve from titration with indicator for part 1 volume of NaOH added mL  $pH$  of the solution

Volume of NaOH added (mL)	$pH$ of the solution
4	6.4
6	4.6
4	6.5
9	5.9
5	9.5
9	5.9

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web background plot of the  $pH$  of a solution against the volume of titrant added is called a titration curve the  $pH$  can be

measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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