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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 = 150.9 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{126.0 \text{ g}}{150.9 \text{ g}} \times 100 = 83.5\%$

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 $80.0 \text{ g } I_2O_5 \cdot \frac{1 \text{ mol } I_2O_5}{283.5 \text{ g } I_2O_5} \cdot \frac{1 \text{ mol } I_2}{2 \text{ mol } I_2O_5} \cdot \frac{253.8 \text{ g } I_2}{1 \text{ mol } I_2} = 28.0 \text{ g } I_2$

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