

MATTER AND ITS CHANGES

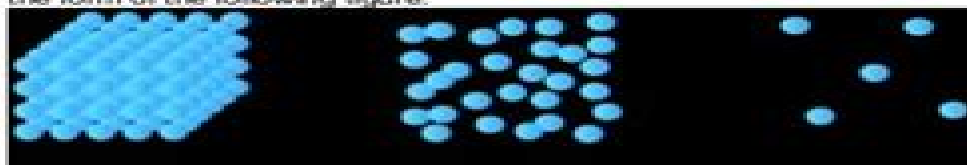
•EFFECT OF CHANGE IN TEMPERATURE ON THE STATES OF MATTER

We know that matter exists in three different physical states – solid, liquid, and gas. The states of matter are interchangeable i.e., a substance can be changed from its solid state to its liquid state and vice-versa. Transitions from the liquid state to the gaseous state and vice-versa are also possible. However, some substances directly change into their gaseous states from their solid states and vice-versa.

Water exists in all three states (i.e., solid, liquid, and gaseous) as ice, water, and water vapour respectively. These states are interchangeable (see the given figure).



Temperature is a major factor responsible for causing a change in the various states of matter. The changes in the states of matter with increasing or decreasing temperatures can be shown in the form of the following figure.



Effect of temperature on the states of matter

On heating a solid substance, the kinetic energy of its constituent particles increases. As a result, the particles start vibrating with greater speed. This extra energy helps the particles to overcome the inter-particle force of attraction. Soon, they leave their positions and start moving more freely. Consequently, the substance melts. The change of solid state into liquid state is known as fusion. The minimum temperature at which a substance starts melting and becomes a liquid at atmospheric pressure is known as its melting point.

It has been observed that temperature remains constant till all solid melts into liquid. The heat supplied is used for changing a solid state into its liquid state by overcoming the particle-particle attraction force. The solid substance absorbs heat energy without showing any rise in temperature. This heat energy is called latent heat. The amount of heat required to convert a unit mass of a solid into its liquid state, without a change in temperature (i.e., at its melting point), is called the latent heat of fusion.

For ice, the latent heat of fusion is 334 kJ/kg . This means that 334 kJ of heat is required to convert 1 kg of ice at 0°C into 1 kg of water at 0°C . Conversely, 334 kJ of heat is released when 1 kg of water freezes at 0°C to give 1 kg of ice at 0°C .

If heating is continued, then the kinetic energy of liquid particles increases further. This increases the velocity of the particles, and at a certain temperature, they obtain enough energy to break free from the force of attraction of the other particles. At this point, the liquid starts boiling i.e., starts

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**James E. Brady, Frederick Senese, Neil
D. Jespersen**



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Matter And Its Changes Gr. 4-6 Doug Sylvester,1997-01-01 In this fast paced unit students discover that matter matters An engaging array of activities combined with interesting worksheets compliments the concepts brought forward in the student notes Relating the study of matter atoms and molecules to the real world is essential Students delight as they learn about DNA fingerprinting and why a grade two class eating pop and chocolate bars is important to the study of chemistry Optional activities add flexibility and an element of fun to the unit Finally a lesson plan on atoms and molecules that will not give students that glazed eye dead fish look This Physical Science lesson provides a teacher and student section with a variety of reading passages activities crossword word search and answer key to create a well rounded lesson plan *Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg,2015-01-16 *Ebook Chemistry The Molecular Nature of Matter and Change* Matter Tom DeRosa,Carolyn Reeves,2009 SCIENCE IS A GREAT AREA TO TEACH BECAUSE CHILDREN HAVE A NATURAL CURIOSITY ABOUT THE WORLD THEY WANT TO KNOW WHY AND HOW THINGS WORK WHAT THINGS ARE MADE OF AND WHERE THEY CAME FROM Hypothesis on MATTER Nainan K. Varghese,2008-05-09 Hypothesis on MATTER is a revolutionary alternative concept which attempts to explain all physical phenomena related to matter based on just one type of fundamental particle the Quantum of matter These particles form what the author calls 2D energy fields Space is assumed to contain an infinite number of 2D Energy Fields extending in all directions Nainan masterfully explains a wide array of physical phenomena from the origin of matter to gravity and subatomic interactions to cosmological events based on the simple mechanical interactions of quanta of matter There is no more any need to envisage actions at a distance or to invoke irrational assumptions like diversity of forces mass energy equivalence constancy of light s speed dual nature of electric charge singularities big bang etc This new concept will radically alter our understanding of the physical universe and at the same time explain complex physical phenomena with simple Cause and Effect relationships **Chemistry** James E. Brady,Frederick Senese,Neil D. Jespersen,2009 **The Mechanics' Magazine** ,1855 The Outline of Man's Knowledge Clement Wood,1927 *Matter* Tom DeRosa,Carolyn Reeves,2009-05 SCIENCE IS A GREAT AREA TO TEACH BECAUSE CHILDREN HAVE A NATURAL CURIOSITY ABOUT THE WORLD THEY WANT TO KNOW WHY AND HOW THINGS WORK WHAT THINGS ARE MADE OF AND WHERE THEY CAME FROM **The Nature of Matter** Anna Claybourne,2007 Discusses matter from atoms and elements and how they bond to the properties of solids liquids and gases *Physical vs. Chemical Changes in Matter | Understanding Chemical Properties of Matter | Grade 6-8 Physical Science* Baby Professor,2024-04-15 Dive into the transformative world of matter with Physical vs Chemical Changes in Matter This essential guide for grades 6 8 demystifies the complexities behind physical and chemical changes It lays a solid foundation for understanding chemical properties from melting ice to rusting metal A must have for educators homeschooling parents and school libraries it s aligned with the US STEM curriculum ensuring young scientists

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