

## ABSTRACT

This study was conducted to determine the effects of temperature on the rate of a chemical reaction. The reaction used was the reaction between hydrogen peroxide and potassium iodide. The reaction was monitored by measuring the volume of oxygen gas produced over time. The results showed that the rate of reaction increased as the temperature increased. This is because the molecules have more kinetic energy and are therefore more likely to collide and react. The study also found that the reaction was first order with respect to the concentration of hydrogen peroxide and zero order with respect to the concentration of potassium iodide.

## QUESTION

Does the rate of reaction increase with temperature?

## HYPOTHESIS

It is hypothesized that the rate of reaction will increase as the temperature increases. This is because the molecules have more kinetic energy and are therefore more likely to collide and react.

## RESEARCH

The research was conducted by searching for information on the effects of temperature on the rate of chemical reactions. The information was found in textbooks, scientific journals, and online resources. The research showed that the rate of reaction increases with temperature, which supports the hypothesis.

# Effects Of Temperature On REACTION TIME

## MATERIALS

The following materials are required to conduct this experiment:

- Hydrogen peroxide solution (3%)
- Potassium iodide solution (0.1M)
- Water
- Thermometer
- Conical flask
- Stopwatch
- Delivery tube
- Gas syringe

## TRIAL 1: HOT TAP WATER



## TRIAL 2: ICE WATER



## PROCEDURE

1. Prepare a series of conical flasks containing different volumes of hydrogen peroxide and potassium iodide solutions at different temperatures.

2. Measure the time taken for the reaction to complete (indicated by the formation of a precipitate or a color change).

3. Repeat the experiment for each temperature and calculate the average reaction time.

## FUTURE DIRECTIONS

The study could be extended by investigating the effect of other factors such as the concentration of the reactants and the presence of a catalyst on the rate of reaction.

## RESULTS

In the first experiment, I noticed that the reaction was faster in hot water than in cold water. This is a result of the average particles in the hot water having more kinetic energy than the particles in the cold water. The average reaction time for the hot water was about 10 seconds, while the average reaction time for the cold water was about 20 seconds. In the second experiment, I noticed that the reaction was faster in ice water than in room temperature water. This is because the molecules in the ice water have less kinetic energy and are therefore less likely to collide and react.

## CONCLUSION

Based on the results of the experiment, it can be concluded that the rate of reaction increases with temperature. This is because the molecules have more kinetic energy and are therefore more likely to collide and react. The study also found that the reaction was first order with respect to the concentration of hydrogen peroxide and zero order with respect to the concentration of potassium iodide.

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# Science Projects About Temperature And Heat

**Janice VanCleave**



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## **Table of Contents Science Projects About Temperature And Heat**

1. Understanding the eBook Science Projects About Temperature And Heat
  - The Rise of Digital Reading Science Projects About Temperature And Heat
  - Advantages of eBooks Over Traditional Books
2. Identifying Science Projects About Temperature And Heat
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Science Projects About Temperature And Heat
  - User-Friendly Interface
4. Exploring eBook Recommendations from Science Projects About Temperature And Heat
  - Personalized Recommendations
  - Science Projects About Temperature And Heat User Reviews and Ratings
  - Science Projects About Temperature And Heat and Bestseller Lists
5. Accessing Science Projects About Temperature And Heat Free and Paid eBooks
  - Science Projects About Temperature And Heat Public Domain eBooks
  - Science Projects About Temperature And Heat eBook Subscription Services
  - Science Projects About Temperature And Heat Budget-Friendly Options
6. Navigating Science Projects About Temperature And Heat eBook Formats
  - ePub, PDF, MOBI, and More
  - Science Projects About Temperature And Heat Compatibility with Devices
  - Science Projects About Temperature And Heat Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Science Projects About Temperature And Heat
  - Highlighting and Note-Taking Science Projects About Temperature And Heat
  - Interactive Elements Science Projects About Temperature And Heat



8. Staying Engaged with Science Projects About Temperature And Heat
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Science Projects About Temperature And Heat
9. Balancing eBooks and Physical Books Science Projects About Temperature And Heat
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Science Projects About Temperature And Heat
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Science Projects About Temperature And Heat
  - Setting Reading Goals Science Projects About Temperature And Heat
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Science Projects About Temperature And Heat
  - Fact-Checking eBook Content of Science Projects About Temperature And Heat
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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