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# School Technology Funding Directory The K12 Decision Makers Guide To Federal And Private Funds 1999

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web thus for reaction 9 2 in which 1 mole of  $H_2SO_4$  and 1 mole of  $NaOH$  combine to form 2 moles of  $H_2O$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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