

Table 1 Li-Rb Crystal Structure and Lattice Parameter Data

Phase	Composition, at.% Rb	Pearson symbol	Space group	Strukturbericht designation	Prototype	Lattice parameters, nm		Comment
						a	c	
(Li).....	~ 0	AP2	<i>P6₃/mmc</i>	A3	Mg	0.3111	0.5093	<75 K
(βLi).....	~ 0	<i>cI2</i>	<i>Im3m</i>	A2	W	0.35093	...	25 °C
(Rb).....	~ 100	<i>cI2</i>	<i>Im3m</i>	A2	W	0.5703	...	25 °C

and they have similar large atomic radii (0.2806 and 0.3024 nm) in comparison to Li (0.1728 nm). In addition, they have similar melting points (39.22 and 28.44 °C) and similar boiling points (688 and 697 °C).

The Li-Rb phase diagram of Fig. 1 was calculated by assuming negligible mutual solubility in the liquid phase. Transition temperatures were taken from the compilations of [83Cha], [King1], and [King2]. Equilibria involving the gas phase were calculated by assuming an ideal gas with $P_{\text{total}} = P_{\text{Li}} + P_{\text{Li}_2} + P_{\text{Cs}} + P_{\text{Cs}_2}} = 1$ bar in equilibrium with the liquid phase, with $a_{\text{Li}}[\text{L}] = 1$. Thermodynamic data for the gas phase calculation were taken from the compilation of

[73Bar]. The small solubility of Rb in liquid Li was estimated by analogy with the Li-Cs system.

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Li-Rb evaluation contributed by C.W. Bale, Centre de Recherche en Calcul Thermochimique, Ecole Polytechnique, Campus de l'Université de Montréal, P.O. Box 6079, Station A, Montréal, Québec, Canada H3C 3A7. This work was partially supported by the United States Department of Energy funds through the Joint Program on Critical Compilation of Physical and Chemical Data coordinated through the Office of Standard Reference Data, National Institute of Standards and Technology. Literature searched through 1988. Professor Bale is the ASM/NIST Data Program Co-Category Editor for binary alkali metal systems.

The K-Zr (Potassium-Zirconium) System

By C.W. Bale
Ecole Polytechnique de Montréal

Equilibrium Diagram

The assessed K-Zr phase diagram is shown in Fig. 1. The system is almost completely immiscible in both the solid and liquid states. There is no evidence of intermetallic formation in this system, nor in any of the other alkali metal-group IV A systems. The extent of the solubility is very limited, and it follows from thermodynamic considerations that the univariant temperatures in Fig. 1 are virtually identical with phase transitions of the pure components. The crystal

structure data for the elements according to [King2] and [King3] are summarized in Table 1.

[70Ste] attempted to measure the solubility of Zr in liquid K in the temperature range 747 to 1068 °C. The study was part of a program to measure the solubilities in liquid K of five transition elements (Mo, W, V, Ti, and Zr). K containing less than 20 wt.% ppm of oxygen was equilibrated in Zr-sealed capsules (~3 mL) for 24 h. The liquid was drained and quenched. The Zr content

Table 1 K-Zr Crystal Structure and Lattice Parameter Data

Phase	Composition, at.% Zr	Pearson symbol	Space group	Strukturbericht designation	Prototype	Lattice parameters, nm		Comment
						a	c	
(K).....	0.0	<i>cI2</i>	<i>Im3m</i>	A2	W	0.5321	...	RT
(αZr).....	100	AP2	<i>P6₃/mmc</i>	A3	Mg	0.32317	0.51476	RT
(βZr).....	100	<i>cI2</i>	<i>Im3m</i>	A2	W	0.3609	...	> 865 °C

Potassium To Zirconium

**Daniel Harrington, George Watkin
Evans, Harry H. Hill, John Paul
Bonardi, John Wesley Marden, Roy O.
Neal, Dean Ernest Woodward, George
St. John Perrott, M. N. Rich**

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