

CHEMISTRY REVISION NOTES

CONSERVATION OF ENERGY

Conservation of energy is the principle that the total energy of a system is constant.

The total energy of a system is the sum of the kinetic energy and the potential energy of the system.

Energy is a scalar quantity. It can be positive or negative. It is the difference between the final energy and the initial energy. If the final energy is greater than the initial energy, the energy is positive. If the final energy is less than the initial energy, the energy is negative.

Energy is conserved in all processes.

If light energy is converted into heat energy, the total energy is conserved. The energy is not lost, it is just converted into a different form.

The energy is conserved in all processes.

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chemical reactions

A chemical reaction is a process in which one or more substances are converted into one or more different substances.

Chemical reactions are classified into two types: exothermic and endothermic.

Exothermic reactions are those in which energy is released. Endothermic reactions are those in which energy is absorbed.

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Energy is conserved in all processes.

PHASES OF MATTER

Phases of matter are solid, liquid, and gas.

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Léo Pomar



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