

reaction for the reaction of copper with silver(I) nitrate?

- A. $\text{Cu} + \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$
- B. $\text{Cu} + 2 \text{Ag}^+ + 2 \text{NO}_3^- \rightarrow \text{Cu}^{2+} + 2 \text{NO}_3^- + 2 \text{Ag}$
- C. $2 \text{Ag}^+ + 2 \text{NO}_3^- \rightarrow 2 \text{NO}_3^- + 2 \text{Ag}^+$
- D. $\text{Cu} + 2 \text{Ag}^+ \rightarrow \text{Cu}^{2+} + 2 \text{Ag}$

7. One way to test for the presence of iron in solution is by adding potassium thiocyanate to the solution. The product when this reagent reacts with iron is FeSCN^{2+} , which creates a dark red color in solution via the following net ionic equation:



How many grams of iron sulfate would be needed to produce 2 moles of FeSCN^{2+} ?

- A. 110 g
- B. 220 g
- C. 400 g
- D. 500 g

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