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$n$  and  $n \times m$  respectively by the law of matrix multiplication their elements are  $e_{ij} \times n \times k \times 1$  a  $k \times j$  d  $j \times a \times j$  d  $j \times j$  and  $f_{ij} \times n \times k \times 1$   $k \times d \times i$  b  $k \times j$  d  $i \times b$   $j$  thus post multiplying a by d is the column operation

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$$A = \begin{pmatrix} 2 & 2 & 0 & 1 & 1 & 0 & 1 & 1 & 0 & 3 \\ 1 & 1 & 3 & 0 & 1 & 1 & 1 & 1 & 1 & 1 \end{pmatrix}$$
 right a  $2 \times 10$  matrix 
$$B = \begin{pmatrix} 2 & 0 & 1 & 1 & 2 & 1 & 1 & 1 & 0 & 1 & 3 & 1 & 1 & 0 & 0 & 1 & 1 & 3 & 1 & 1 & 5 \end{pmatrix}$$
  $5 \times 20$  problem 2

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web evaluate the determinant of a 2x2x2x2 matrix if a matrix has the same number of rows and columns we call it a square matrix each square matrix has a real number associated with it called its determinant to find the determinant of the square matrix  $\begin{pmatrix} a & b & c & d \\ a & b & c & d \end{pmatrix}$  we first write it as  $\begin{pmatrix} a & b & c & d \\ a & b & c & d \end{pmatrix}$

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web so there are two quantities which can be found in a matrix minors and cofactors minors exist for each element of the matrix and are defined as the determinant of the matrix you get when you ignore the row and column of the element in

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web the above equation works only for neutralizations in which there is a 1:1 ratio between the acid and the base the

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web the titration of a weak acid with a strong base involves the direct transfer of protons from the weak acid to the hydroxide ion the reaction of the weak acid acetic acid with a strong base naoh can be seen below in the reaction the

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web as seen in the chapter on the stoichiometry of chemical reactions titrations can be used to quantitatively analyze solutions for their acid or base concentrations in this section we will explore the underlying chemical equilibria that make acid base titrimetry a

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